Evaluate the experiments that led to the conclusion that electrons are negatively charged particles found in all matter.

The deflection toward positively charged plates demonstrated the negatively charged nature of electrons; the fact that changing the type of electrode or the type of gas used in the cathode ray tube did not affect the ray produced led to the conclusion that electrons are present in all matter.

Compare the relative charge and mass of each of the subatomic particles.

Particle	Relative charge	Relative mass	
Electroin	-1	1/1840	
Proton	+1	100	
Neutron	0	65-16	

11. Calculate What is the difference expressed in kilograms between the mass of a proton and the mass of an electron?

 $1.673 \times 10^{-28} \,\mathrm{g} - 9.11 \times 10^{-28} \,\mathrm{g} = 1.672 \times 10^{-24} \,\mathrm{g} \times 1 \,\mathrm{kg/1000} \,\mathrm{g} = 1.672 \times 10^{-27} \,\mathrm{kg}$

Section 4.3 How Atoms Differ

pages 115-121

Practice Problems

pages 116-121

12. How many protons and electrons are in each atom?

a. radon, 86 protons and 86 electrons

b. magnesium, 12 protons and 12 electrons

13. An atom of an element contains 66 electrons. Which element is it?

dysprosium

14. An atom of an element contains 14 protons. Which element is it?

silicon

15. Challenge Do the atoms shown in the figure to the right have the same atomic number?



Yes. 9

16. Determine the number of protons, electrons, and neutrons for isotopes b through f. in the table above. Name each isotope, and write its symbol.

	Protons and Electrons	Neutrons	Isotope	Symbol
đ	20	26	calcium-46	46 Ca 20 Ca
1	8	9	oxygen-17	120
	26	31	iron-57	∰Fe.
	30	34	zinc-64	\$\$Zm
	80	124	mercury-204	^{20th} Hg

17. Challenge An atom has a mass number of 55. Its number of neutrons is the sum of its atomic number and 5. How many protons, neutrons, and electrons does this atom have? What is the identity of this atom?

25 protons, 25 electrons, 30 neutrons, manganese

18. Boron (B) has two naturally occurring isotopes: boron-10 (abundance = 19.8%, mass = 10.013 annu), boron-11 (abundance = 80.2%, mass = 11.009 annu). Calculate the atomic mass of boron.

B-10 10.013 amu × 0.198 = 1.98 amu

B-11 11,009 amu × 0,802 = 8,83 amu

Atomic mass 1.98 amu + 8.83 amu = 10.81 amu

19. Challenge Nitrogen has two naturally occurring isotopes, N-14 and N-15. Its atomic mass is 14.007. Which isotope is more abundant? Explain your answer.

N-14 is more abundant because the atomic mass is closer to 14 than 15.

Chapter 4 Atomic Structure Test Answers 2012

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and management of common disorders Resources for providers and families are also included throughout the text for further information Expert editor team is well versed in the scope of practice and knowledge base of Pediatric Nurse Practitioners PNPs and Family Nurse Practitioners FNPs Science Books & Films ,1989 Environmental Law in Bosnia and Herzegovina Dragoljub Todić, Mladenka Ignjatić, Vid Vukasović †, 2022-11-20 Derived from the renowned multi volume International Encyclopaedia of Laws this book provides ready access to legislation and practice concerning the environment in Bosnia Herzegovina A general introduction covers geographic considerations political social and cultural aspects of environmental study the sources and principles of environmental law environmental legislation and the role of public authorities The main body of the book deals first with laws aimed directly at protecting the environment from pollution in specific areas such as air water waste soil noise and radiation Then a section on nature and conservation management covers protection of natural and cultural resources such as monuments landscapes parks and reserves wildlife agriculture forests fish subsoil and minerals Further treatment includes the application of zoning and land use planning rules on liability and administrative and judicial remedies to environmental issues There is also an analysis of the impact of international and regional legislation and treaties on environmental regulation Its succinct yet scholarly nature as well as the practical guality of the information it provides make this book a valuable resource for environmental lawyers handling cases affecting Bosnia Herzegovina Academics and researchers as well as business investors and the various international organizations in the field will welcome this very useful guide and will appreciate its value in the study of comparative environmental law and policy

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