

Multiple choice questions 1-10: circle the correct answer(A-E)

Show your work questions 11-14: fill in the answer

1. Which of the following pairs are both electrolytes?

- A. NaOH and SO₂ **B. H₂SO₄ and NH₄Cl** C. CH₃OH and C₂H₅OH D. HNO₂ and NO₂ E. CO₂ and H₂CO₃

2. Which of the following is a weak electrolyte?

- A. HBr(aq) B. KOH(aq) C. HNO₃ **D. H₃PO₄** E. Ba(OH)₂

3. For which of the following pairs are both compounds insoluble in water?

- A. NaBr, Na₂SO₄ B. Na₂CO₃, CaCO₃ **C. Fe₂(CO₃)₃, Ba₃(PO₄)₂** D. Na₃PO₄, Ca₃(PO₄)₂ E. PbSO₄, MgSO₄

4. Which of the following is a weak acid?

- A. HClO₃ B. HCl C. HClO₄ D. HNO₃ **E. HF**

5. Which of the following is a Bronsted acid?

- A. NH₄⁺(aq)** B. CH₃COO⁻(aq) C. NH₃(aq) D. OH⁻(aq) E. Na⁺(aq)

6. Which of the following is a Bronsted base?

- A. NH₄⁺(aq) B. Na⁺(aq) C. HNO₃(aq) **D. NH₃(aq)** E. CH₃COOH(aq)

7. In the reaction: $\text{NH}_3(\text{aq}) + \text{H}_2\text{O} \rightarrow \text{NH}_4^+(\text{aq}) + \text{OH}^-$

- A. In this reaction H₂O is the Bronsted base B. NH₃ is the Bronsted acid and NH₄⁺ is the Bronsted base

- C. OH⁻ is a proton donor D. NH₄⁺ is a Bronsted base

E. NH₃ is the Bronsted acid and NH₃ is the Bronsted base, they are a conjugate acid-base pair

8. The oxidation numbers of S in SO₃²⁻ and SO₄²⁻ are:

- A. +2 and +4 B. +3 and +6 **C. +4 and +6** D. +6 and +8 E. -4 and -6

9. The oxidation number of Cr in the dichromate ion, Cr₂O₇²⁻ is:

- A. +6** B. +5 C. +4 D. -4 E. +7

10. In the reaction: $\text{FeO} + \text{CO} \rightarrow \text{Fe} + \text{CO}_2$

- A. Fe is the oxidizing agent B. CO is reduced C. O²⁻ is oxidized **D. CO is the reducing agent**

Look at the oxidation numbers; O is -2 in all compounds, no change CO is not reduced, it is oxidized; C goes from +2 to +4; The Fe(II) is reduced to the element Fe, so CO is the reducing agent, it reduces the Fe²⁺ and in the process is oxidized itself.

Fill in the answer for the following questions.

11. Chang 4.47 (with some changes).
the following compounds and ions:

Give the oxidation number of the underlined> element in

<u>C</u> O ₂	<u>C</u> H ₄	<u>C</u> ₂ H ₂	<u>C</u> rO ₃	K ₂ <u>C</u> rO ₄	NaH <u>C</u> O ₃	<u>P</u> O ₄ ³⁻
Ox.#: +4	-4	-1	+6	+6	+4	+5
					Na +1, H +1, O -2	

Chapter 4 Reactions In Aqueous Solutions Answers

Terry C. Jones



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