

## Flame Test Lab

### Activity Key

#### Background

The heat from a laboratory burner will cause the ions of some elements to give off light. Electrons will absorb the heat energy from the flame and will “jump” to a higher energy level. When the electrons return to their original energy levels, this absorbed energy is released as light. Different elements absorb and release different amounts of energy and thus produce different kinds of light. This light may be visible or invisible, infrared or ultraviolet radiation. Analysis of this light can be used to identify elements and molecules. In today’s lab, we will observe the visible range of light emission.

#### Objectives

1. Perform a flame test to identify the characteristic color of metal ions.
2. Identify the metal ion in an unknown solution.
3. Calculate the energy of emitted photons.

#### Materials: (per lab group)

10 ml dilute solutions of the following:

Barium Chloride  
Calcium Chloride  
Lithium Chloride  
Potassium Chloride  
Sodium Chloride  
Strontium Chloride  
Unknown Solution

7 Wood Splints  
1 Bunsen Burner  
Safety Goggles & Aprons

#### Note:

If chloride compounds are not available, metal nitrate compounds may be substituted.

Use dilute or approximately 0.1 M solutions

#### Unknowns:

Number the beakers 1-12.  
Add solutions as follows:

- |                       |                        |
|-----------------------|------------------------|
| 1. Strontium Chloride | 7. Calcium Chloride    |
| 2. Calcium Chloride   | 8. Strontium Chloride  |
| 3. Potassium Chloride | 9. Barium Chloride     |
| 4. Barium Chloride    | 10. Potassium Chloride |
| 5. Lithium Chloride   | 11. Sodium Chloride    |
| 6. Sodium Chloride    | 12. Lithium Chloride   |

#### Procedure

1. Put on your safety goggles and an apron.
2. Obtain a wood splint and the barium chloride solution.
3. Dip the wood splint into the solution and allow it to soak for 1-2 minutes.
4. Light Bunsen Burner. USE CAUTION!
5. Remove the splint from the solution and hold the end of the splint into the flame. Do not burn the splint.
6. Observe and record the color of the flame.
7. Repeat steps 2-6 using the remaining chloride solutions.
8. Obtain an unknown solution from your teacher. Record the number of your unknown.
9. Soak a wood splint in the unknown solution.
10. Hold the splint in the flame and record the color of the flame that is produced.
11. Using your data, identify the metal ion in your unknown solution.

# Chapter 5 Flame Tests For Metals Worksheet Answers

**Rex M. Heyworth**



## **Chapter 5 Flame Tests For Metals Worksheet Answers:**

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explanation

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aq c3h7no2 aq h3o aq c3h8no 2 aq h2o l ch3o2h aq nh3 aq ch3co 2 aq nh 4 aq sbf5 aq 2hf aq h2f aq sbf 6 aq identify the

conjugate acid base

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web jul 22 2023 correct answer c kh2po4 explanation kh2po4 is an acid salt because it is derived from a weak acid

phosphoric acid and a strong base potassium hydroxide it contains the hydrogen ion from the acid making it acidic but also has the ability to form basic ions due to the presence of the potassium ion

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635 640 write balanced equations for reactions between the following aluminum and sulfuric acid 2al s 1 3h 2so 4 aq 0 al 2

so 4 3 aq 1 3h 2 g calcium carbonate and hydrobromic acid

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each other acids mixed with bases neutralize each other acids mixed with bases make stronger bases acids mixed with bases



make stronger acids acids and bases interact with each other in what is called a neutralization reaction

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arrhenius base a compound that increases the concentration of hydroxide ion  $\text{OH}^-$  in aqueous solution the reaction of an acid

and a base exercises give two examples of arrhenius acids give two examples of arrhenius

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web 1 pt during science lab students were asked to test a variety of common household substances and classify them as acid

base or neutral tests on shampoo produced the following results felt slippery  $\text{pH}$  8.5 turned red litmus blue these results

indicate shampoo is an acid shampoo is a base

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web nov 13 2022 suggest simple tests you could carry out to determine if an unknown substance is an acid or a base state

the chemical definitions of an acid and a base in terms of their behavior in water write the formula of the salt formed when a

given acid and base are combined

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web the ionization of weak acids and weak bases many acids and bases are weak that is they do not ionize fully in aqueous

solution a solution of a weak acid in water is a mixture of the nonionized acid hydronium ion and the conjugate base of the

acid with the nonionized acid present in the greatest concentration

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sound wave and speed characteristics of

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web oct 10 2022 correct answer d salt water explanation neutralization is a chemical reaction between an acid and a base resulting in the formation of a salt and water the product of neutralization is therefore salt water as the acid and base combine to form a salt compound and water molecules rate this question 4