

CHAPTER 5 Electrons in Atoms

- Orange light has a frequency of $4.8 \times 10^{14} \text{ s}^{-1}$. What is the energy of one quantum of orange light?
- Which is greater, the energy of one photon of orange light or the energy of one quantum of radiation having a wavelength of $3.36 \times 10^{-8} \text{ m}$?
- Use the relationships $E = h\nu$ and $c = \lambda\nu$ to write E in terms of h , c , and λ .
- A radio station emits radiation at a wavelength of 2.90 m. What is the station's frequency in megahertz?
- Record the frequency of your favorite radio station. What is the wavelength of the radiation emitted from the station?
- List the sequence in which the following orbitals fill up: 1s, 2s, 3s, 4s, 5s, 6s, 7s, 2p, 3p, 4p, 5p, 6p, 7p, 3d, 4d, 5d, 6d, 4f, 5f.
- Which element has the ground-state electron configuration $[\text{Kr}]5s^2 4d^{10} 5p^4$?
- Which element has the ground-state electron configuration $[\text{Ar}]4s^2 3d^{10}$?
- Write electron-dot structures for the following atoms.
 - $[\text{Ne}]3s^2 3p^3$
 - $[\text{Ar}]4s^2 3d^3$
 - potassium

10. Complete the following table.

Element	Symbol	Orbitals					Electron Configuration
		1s	2s	2p _x	2p _y	2p _z	
a. Nitrogen							$1s^2 2s^2 2p^3$
b.	F	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	\uparrow	
c. Carbon							
d.							$1s^2 2s^1$

Chapter 5 Supplemental Problems Electrons In Atoms Answers

Wolfson



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