Types of Chemical Reactions Worksheet #1

HI/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 2Mg + O₂ → 2 MgO Synthesis
- Cal₂ + Cl₂ → CaCl₂ + I₂ Single Replacement
- 3) 3KOH + AlCl₃ → Al(OH)₃ + 3KCl Double Replacement
- C + O₂ → CO₂ Synthesis
- *5) Ca(OH)₂ + 2HCl → CaCl₂ + 2H₂O Double Replacement
- 6) 2KClO₃ → 2KCl+3 O₂ Decomposition
- 7) 3CuSO₄ + 2Al → Al₂(SO₄)₃ + 3Cu Single Replacement
- 8) Na₂S + 2AgNO₁ → 2NaNO₁ + Ag₂S Double Replacement
- 9) $3H_2 + N_2 \rightarrow 2NH_3$ Synthesis
- *10) 2Na + 2H2O -> 2NaOH + H2 Single Replacement

*Remember: "HyO" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- Double replacement: Na₂CrO₄ + PbCl₂ → 2 NaCl + PbCrO₄
- Single replacement: Cl₂ + 2 NaBr → 2 NaCl + Br₂
- Decomposition: Mg(ClO₃)₂ → MgCl₂ + 3 O₂
- 14) Synthesis: 2 H₂ + O₂ → 2 H₂O
- 15) Double replacement: 3 Ca(OH)₂ + 2 FeCl₃ → 2 Fe(OH)₁ + 3 CaCl₂
- 16) Single replacement: Fe + Cu(NO₃)₂ → Fo(NO₁)₂ + Cu [Assume Fe²⁺]
- Decomposition: 2 Hg₂O → 4 Hg + O₂
- 18) Synthesis: $S + O_2 \rightarrow SO_2$
- Double replacement: AgNO₃ + KI → AgI + KNO₃
- 20) Single replacement: Cu + 2 AgNO₃ → 2 Ag + Cu(NO₃)₂ [Copper (II) is used here]

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K. S. Verma

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