Brief Instructions

An electron configuration is a method of indicating the arrangement of electrons about a nucleus. A typical electron configuration consists of numbers, letters, and superscripts with the following format:

- A number indicates the energy level (The number is called the principal quantum number.).
- A letter indicates the type of orbital; s, p, d, f.
- A superscript indicates the number of electrons in the orbital. Example: Is² means that
 there are two electrons in the 's' orbital of the first energy level. The element is helium.
 How To write an electron configuration:
- A. Determine the total number of electrons to be represented.
- B. Use the Aufbau process to fill the orbitals with electrons. The Aufbau process requires that electrons fill the lowest energy orbitals first. In another words, atoms are built from the ground upwards.
- C. The sum of the superscripts should equal the total number of electrons. Example: 12Mg Is² 2s² 2p⁶ 3s²

Configuration Writing Practice

Write a **ground state** electron configuration for each neutral atom. **Ground state** means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

| -85 | 1. | Na | 4. | U | 7. | Ti | 10. | H |
|-----|----|----|----|----|----|----|-----|---|
| | 2. | Pb | 5. | N | 8. | Ce | g | |
| | 3. | Sr | 6. | Ag | 9. | CI | | |

Write a ground state electron configuration for these ions. Remember that ions have a change in the total number of electrons (positive have lost electrons and negative have gained). Example: N³ is 1s² 2s² 2p⁵. It has three extra electrons

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|---------------------------------|--|--|--|--|--|
| 11. | 12. | 13. | 14. | 15. | 16. |
| O-2 | Fe ²⁺ | B3+ | Ni ²⁺ | K* | Co3+ |
| 19.70 | . If words weblind a | we heald a executions | on of two alacteans. | The many than many a soll of | the effective and the second |

17. If each orbital can hold a maximum of two electrons, how many electrons can each of the following hold?

| a. | 2s | c. 4f | e. 4d |
|----|----|-------|-------|
| b. | 5p | d. 3d | |

18) Isoelectronic species have similar electron configurations. Which of these are isoelectronic?

a) Li⁺, H⁻, He
b) Ca²⁺, Ne, S²
For the following electron configurations choose 3 possible elements (or ions) they may represent

- 19) 1s22s22p63s23p64s23d264p4
- 20) 1s² 2s²2p⁶ 3s²3p⁶ 4s² 3d¹⁰ 4p⁵
- 21) 1s22s22p63s23p64s23d104p6
- 22) 1s² 2s²2p⁶ 3s²3p⁶ 4s² 3d¹⁰ 4p⁶ 5s² 4d¹⁰ 5p⁶ 6s² 4f¹⁴ 5d¹⁰ 6p⁶ 7s¹
- 23) 1s² 2s²2p⁶ 3s²3p⁶ 4s² 3d¹⁰ 4p⁶ 5s²4 d¹⁰ 5p⁸ 6s² 4f¹⁴ 5d¹⁰ 6p⁶ 7s² 5f¹⁴ 6d⁸
- 24) 1s² 2s²2p⁶ 3s²3p⁶ 4s² 3d¹⁰ 4p⁶ 5s² 4d¹⁰ 5p⁶ 6s² 4f¹⁰
- 25) 1s² 2s²2p6 3s²3p6 4s² 3d¹0 4p6 5s² 4d¹0 5p6 6s² 4f²4 5d¹0 6p4
- 26) 1s² 2s²2p6 3s²3p6 4s² 3d³
- 27) 1s² 2s²2p⁶ 3s²3p⁶ 4s² 3d¹⁰ 4p⁶ 5s² 4d¹⁰ 5p²
- 28) [Kr] 5s² 4d¹⁰ 5p³
- 29) [Kr] 5s² 4d¹⁰ 5p⁶
- 30) [Ar] 4s1
- 31) [Xe] 6s2 4f10
- 32) [Xe] 6s2 4f14 5d7
- 33) [Ne] 3s² 3p¹

Brief Instructions Configurations Writing Practice Answer

Robert W. Barnett, Jacob S. Blumner

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