

CHEMICAL BONDING

Type of bonds

- Chemical bond formation takes place when
(A) Energy is absorbed
(B) Forces of attraction are overcome by the forces of repulsion
(C) Forces of repulsion are overcome by forces of attraction
(D) Forces of attraction are equal to the forces of repulsion
- Most favourable conditions for ionic bond formation are
(A) Low charge on ions, large cation, small anion
(B) High charge on ions, large cation, small anion
(C) High charge on ions, small cation, large anion
(D) High charge on ions, small cation, large anion
- Which of the following shows the higher lattice energy?
(A) RbF (B) CsF (C) NaF (D) KF
- Which of the following is least soluble in water?
(A) BaF₂ (B) SiF₄ (C) CaF₂ (D) MgF₂
- Which of the following sets does not contain any ionic species:
(A) NH₄Cl, OF₂, H₂S (B) CO₂, Cl₂, CCl₄ (C) BF₃, AlF₃, TiF₃ (D) I₂, CaO, CH₃Cl
- There is no S-S bond in :
(A) S₂O₃²⁻ (B) S₂O₄²⁻ (C) S₄O₆²⁻ (D) S₂O₇²⁻
- The species that does not contain peroxide ions are :
(A) PbO₂ (B) H₂O₂ (C) SrO₂ (D) BaO₂
- Which of the following compounds contains electrovalent and covalent bonds?
(A) CS₂ (B) SO₂ (C) KCN (D) Phosphine
- In which solvent has maximum solubility
(A) H₂O (B) C₂H₅OH (C) CH₃COCH₃ (D) C₂H₅OC₂H₅
- Which of the following is a non-directional bond
(A) Metallic bond (B) Covalent bond (C) Co-ordinate bond (D) All are directional
- Find the number of σ and π bonds in the following compounds
(A) CH₃CH₃ (B) C₆(CN)₆ (C) CH₂=CH₂ (D) C₂H₂Cl
- Point out the inter particle force operating in each of the following :
(A) H₂O(l) (B) H₂O(s) (C) Dryice (D) KCl
(E) Hg (F) CF₄ (G) Rock salt (H) Diamond
- True / False
(i) Graphite, a covalent molecule, conducts current in solid state
(ii) Intra molecular H-Bonding decreases the solubility of compound in water

Matching

	Column I		Column - II
(a)	H ₂ O	(p)	Ionic bond
(b)	C ₂ H ₂	(q)	Covalent bond
(c)	KCN	(r)	Co-ordinate bond
(d)	K ₄ [Fe(CN) ₆]	(s)	Hydrogen bond

Fagan's rule, Physical property

- Amongst LiCl, BeCl₂, MgCl₂ and RbCl the compounds with greatest and least ionic character respectively are :
(A) LiCl and RbCl (B) RbCl and BeCl₂ (C) RbCl and MgCl₂ (D) MgCl₂ and BeCl₂
- The formal charge of the O atoms in the ion $[\text{O}=\text{N}=\text{O}]_{-2}$ is

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