

# Chapter 10 Solutions Practice Problems

## Solution Concentration

1. a. HCl was collected when burned by adding 20.0 g of 20.0% solution (20.0% means 20.0 g/g total) to 10.0 g of water (H<sub>2</sub>O, water density = 1.00 g/mL, density = 1.00 g/mL).
- b. What is the molarity of this solution?

$$\frac{20.0 \text{ g HCl}}{1} \times \frac{1 \text{ mol}}{36.46 \text{ g}} = 0.549 \text{ mol HCl}$$

$$\frac{20.0 \text{ g HCl}}{1} \times \frac{1 \text{ L}}{1000 \text{ g}} = 0.020 \text{ L} \quad M = \frac{0.549 \text{ mol}}{0.020 \text{ L}} = \boxed{27.5 \text{ M}}$$

- c. What is the mass percent of HCl in solution in this solution?

$$\frac{20.0 \text{ g HCl}}{1} \times \frac{1 \text{ mol}}{36.46 \text{ g}} = 0.549 \text{ mol HCl}$$

$$M_{\text{HCl}} = \frac{0.549 \text{ mol} \times 36.46 \text{ g/mol}}{0.549 \text{ mol} + 0.020 \text{ L} \times 1000 \text{ g/L}} = \boxed{0.101}$$

$$\begin{aligned} M_{\text{HCl}} &= 1 \times M_{\text{HCl}} \\ &= 0.101 \times 100 \\ &= 10.1\% \end{aligned}$$

- d. What is the mass percent of HCl in solution in this solution?

$$\frac{20.0 \text{ g HCl}}{1} \times \frac{1 \text{ mol}}{36.46 \text{ g}} = 0.549 \text{ mol HCl} \quad \frac{0.549 \text{ mol}}{0.020 \text{ L}} = 27.5 \text{ M}$$

- e. What is the molarity of this solution?

$$\frac{20.0 \text{ g HCl}}{1} \times \frac{1 \text{ mol}}{36.46 \text{ g}} = 0.549 \text{ mol HCl}$$

$$\frac{0.549 \text{ mol} \times 36.46 \text{ g/mol}}{0.020 \text{ L} \times 1000 \text{ g/L}} = \boxed{1.01 \text{ M}}$$

# Chemistry Ch 18 Solutions Practice Problems Answers

**D Kirk**



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Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science      **Ebook:**

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Moore, Peter J. Mikulecky, 2022-11-23 Everything you need to crush chemistry with confidence Chemistry All in One For Dummies arms you with all the no nonsense how to content you'll need to pass your chemistry class with flying colors You'll find tons of practical examples and practice problems and you'll get access to an online quiz for every chapter Reinforce the concepts you learn in the classroom and beef up your understanding of all the chemistry topics covered in the standard curriculum Prepping for the AP Chemistry exam Dummies has your back with plenty of review before test day With clear definitions concise explanations and plenty of helpful information on everything from matter and molecules to moles and measurements Chemistry All in One For Dummies is a one stop resource for chem students of all valences Review all the topics covered in a full year high school chemistry course or one semester of college chemistry Understand atoms molecules and the periodic table of elements Master chemical equations solutions and states of matter Complete practice problems and end of chapter quizzes online Chemistry All In One For Dummies is perfect for students who need help with coursework or want to cram extra hard to ace that chem test      **Student's Guide to Chemistry; a Modern Introduction** David

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**Chemistry** Hall,1995-12 Fundamentals of Chemistry David E. Goldberg,2001 Chemistry John Christian Bailar,1984 The Athenaeum ,1872

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