

- The amount of a gas in a mole is directly related to its volume at constant pressure and temperature.
- The properties of gases can be varied. The relationships of these properties can be quantified experimentally with the aid of the different laboratory apparatus or by using the different gas laws as follows:

• Boyle's Law	$V_1 P_1 = V_2 P_2$
• Charles' Law	$V_1 / T_1 = V_2 / T_2$
• Gay-Lussac's Law	$P_1 / T_1 = P_2 / T_2$
• Avogadro's Law	$V_1 n_1 = V_2 n_2$
• Combined Gas Law	$V_1 P_1 / T_1 = V_2 P_2 / T_2$
• Ideal Gas Law	$PV = nRT$

Where V = volume,

P = pressure,

T = temperature,

n = amount of gas in moles,

$R = 0.0821 \text{ L.atm./mol.K}$

1 = used as a subscript means: initial condition

2 = used as a subscript means: final condition

- Not all gases behave ideally. Most of the gases found in nature conform to the principles of Boyle's Law, Charles' Law, Gay-Lussac's Law, Avogadro's Law, and Combined Gas Law.
- The following conversion factors are useful in solving gas law related problems:

i. For volume

$$1 \text{ mL} = 1 \text{ cm}^3$$

$$1 \text{ L} = 1 \text{ dm}^3$$

$$1 \text{ m}^3 = 1000 \text{ L}$$

<http://www.metric-conversions.org/volume/cubic-meters-to-liters.htm>

j. For pressure

$$1 \text{ atm} = 760 \text{ mmHg} = 76 \text{ cmHg} = 760 \text{ torr} = 101,325 \text{ Pa} = 14.6956 \text{ psi}$$

k. For temperature

$$0^\circ\text{C} = 273.15\text{K}$$

$$0^\circ\text{C} = 32^\circ\text{F}$$

Chemistry Chapters 12 Gases Solving Problems

David Goldberg



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showcases the rapidly developing field of materials optics It is aimed at a broad audience as the readers require only

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