

# Chemistry

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(Chapter 5) (Coordination Compound) (Intext Questions)  
(Class 12)

## Question 5.5:

Explain on the basis of valence bond theory that  $[\text{Ni}(\text{CN})_4]^{2-}$  ion with square planar structure is diamagnetic and the  $[\text{NiCl}_4]^{2-}$  ion with tetrahedral geometry is paramagnetic.

## Answer 5.5:

Ni is in the +2 oxidation state i.e., in  $d^8$  configuration.

$d^8$  configuration: 

There are 4  $\text{CN}^-$  ions. Thus, it can either have a tetrahedral geometry or square planar geometry. Since  $\text{CN}^-$  ion is a strong field ligand, it causes the pairing of unpaired 3d electrons.



It now undergoes  $dsp^2$  hybridization. Since all electrons are paired, it is diamagnetic. In case of  $[\text{NiCl}_4]^{2-}$ ,  $\text{Cl}^-$  ion is a weak field ligand. Therefore, it does not lead to the pairing of unpaired 3d electrons. Therefore, it undergoes  $sp^3$  hybridization.



Since there are 2 unpaired electrons in this case, it is paramagnetic in nature.

## Question 5.6:

$[\text{NiCl}_4]^{2-}$  is paramagnetic while  $[\text{Ni}(\text{CO})_4]$  is diamagnetic though both are tetrahedral. Why?

## Answer 5.6:

Though both  $[\text{NiCl}_4]^{2-}$  and  $[\text{Ni}(\text{CO})_4]$  are tetrahedral, their magnetic characters are different. This is due to a difference in the nature of ligands.  $\text{Cl}^-$  is a weak field ligand and it does not cause the pairing of unpaired 3d electrons. Hence,  $[\text{NiCl}_4]^{2-}$  is paramagnetic.



In  $\text{Ni}(\text{CO})_4$ , Ni is in the zero oxidation state i.e., it has a configuration of  $3d^8 4s^2$ .



But  $\text{CO}$  is a strong field ligand. Therefore, it causes the pairing of unpaired 3d electrons. Also, it causes the 4s electrons to shift to the 3d orbital, thereby giving rise to  $sp^3$  hybridization. Since no unpaired electrons are present in this case,  $[\text{Ni}(\text{CO})_4]$  is diamagnetic.

## Question 5.7:

$[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$  is strongly paramagnetic whereas  $[\text{Fe}(\text{CN})_6]^{3-}$  is weakly paramagnetic. Explain.

## Answer 5.7:

In both  $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$  and  $[\text{Fe}(\text{CN})_6]^{3-}$ , Fe exists in the +3 oxidation state i.e., in  $d^5$  configuration.

$d^5$  

Since  $\text{CN}^-$  is a strong field ligand, it causes the pairing of unpaired electrons. Therefore, there is only one unpaired electron left in the d-orbital.



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pediatrics yellow tops light blue citrated tubes correct answer tests pt ptt 1nr clotting additives sodium citrate sst tubes  
correct answer there is a gel at the bottom comes in a gold and tiger top always

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collections on hours to a few days old a primary consultant on the timing for collecting blood for drug levels is found in which  
of the following