

Complete the following conversions. Don't forget to use the correct number of significant figures.

$$1.45 \text{ km} = \text{_____ m}$$

$$92.00 \text{ mm} = \text{_____ km}$$

$$0.06 \text{ m} = \text{_____ cm}$$

$$4.11 \times 10^{-4} \text{ km} = \text{_____ in}$$

$$36.4 \text{ ft} = \text{_____ mi}$$

$$0.416 \text{ mi} = \text{_____ cm}$$

$$94.52 \text{ nm} = \text{_____ in}$$

$$88.0 \text{ mg} = \text{_____ g}$$

$$0.620 \text{ kg} = \text{_____ g}$$

$$1.79 \times 10^6 \text{ mg} = \text{_____ kg}$$

$$43.5 \text{ kg} = \text{_____ lb}$$

$$128.0 \text{ oz} = \text{_____ g}$$

$$750 \text{ g} = \text{_____ lb}$$

$$32.8 \text{ L} = \text{_____ mL}$$

$$4.8 \text{ cm}^3 = \text{_____ L}$$

$$7.35 \text{ gal} = \text{_____ L}$$

$$82.40 \text{ mL} = \text{_____ pint}$$

$$100 \text{ qt} = \text{_____ kL}$$

$$2.29 \text{ cm}^3 = \text{_____ in}^3$$

$$1.25 \text{ m}^3 = \text{_____ ft}^3$$

$$0.0236 \text{ oz/hr} = \text{_____ mg/min}$$

# Chemistry Conversion Factors Practice Problems Answers

**Neil D. Jespersen, Pamela Kerrigan**



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