
Bond Energy Study Guide (Chemistry 2.2)

1. Why atoms form bonds

Atoms form bonds to reach a lower, more stable energy state.

- Each atom has **protons (+)** in its nucleus and **electrons (-)** around it.
- **Coulombic attraction** is the force between opposite charges. When atoms get close, their nuclei attract the other atom's electrons.
- This attraction lowers potential energy — that's why bonds form.

When the bond forms:

- **Energy is released** → the system becomes more stable.
- **Bond making = exothermic process.**

When a bond breaks:

- **Energy is absorbed** → you must overcome that attraction.
- **Bond breaking = endothermic process.**

2. Endothermic vs. Exothermic

Type of Process	Energy Flow	What Happens	ΔH (Enthalpy Change) Sign
Endothermic	Energy absorbed	Bonds are broken	Positive (+)
Exothermic	Energy released	Bonds are formed	Negative (-)

Example:

- When hydrogen and chlorine form HCl, the reaction releases energy — **exothermic**.
- When you separate HCl into H and Cl atoms, you need energy — **endothermic**.

3. Net Enthalpy Change (ΔH)

Chemistry Energy Study Guide Key

Brendan G. Carr



Chemistry Energy Study Guide Key:

Chemistry Energy Study Guide Key Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Chemistry Energy Study Guide Key**," written by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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