

## EQUILIBRIUM CONSTANT - WORKSHEET 2\_ANSWERS

1. Write the equilibrium constant expression for each of the following reactions:



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2] [\text{O}_2]}$$



$$K = \frac{[\text{C}_2\text{H}_6]}{[\text{H}_2]^3}$$



$$K = \frac{1}{[\text{I}_2 (\text{aq})]}$$

2. For the following reaction:



At equilibrium,  $[\text{Fe}^{2+}] = 0.50 \text{ M}$ ,  $[\text{Ag}^+] = 1.0 \text{ M}$ ,  $[\text{Fe}^{3+}] = 1.50 \text{ M}$

- a) Calculate the equilibrium constant for the reaction.

$$K = 3.0$$

- b) If at equilibrium,  $[\text{Fe}^{2+}] = 0.20 \text{ M}$  and  $[\text{Ag}^+] = 0.30 \text{ M}$ , what must be the concentration of  $\text{Fe}^{3+}$ ?

$$[\text{Fe}^{3+}] = 0.18 \text{ mol/L}$$

3.  $2 \text{NO} (\text{g}) + \text{O}_2 (\text{g}) \rightleftharpoons 2 \text{NO}_2 (\text{g}) \quad K = 5.0 \times 10^{-5}$

Calculate the equilibrium concentration of oxygen gas if the NO and  $\text{NO}_2$  concentrations are equal at equilibrium.

$$[\text{O}_2] = 2.0 \times 10^{-6} \text{ mol/L}$$

4. a)  $\text{CH}_3\text{COOH} (\text{aq}) \rightleftharpoons \text{H}^+ (\text{aq}) + \text{CH}_3\text{COO}^- (\text{aq}) \quad K = 1.8 \times 10^{-5}$



Which of the above reactions would have the greatest amount of products present at equilibrium compared to reactants?

Reaction b. Since K is larger the amount of products must be larger, because K is products/reactants.

5. For the reaction:



at an equilibrium temperature of  $1000^\circ\text{C}$ , a 1.00 L flask contains 0.120 moles of ammonia, 1.03 moles of nitrogen and 1.62 moles of hydrogen. Calculate the equilibrium constant at this temperature.

$$K = 3.29 \times 10^{-2}$$

# Chemistry Equilibrium Practice Problems With Answers

**Darrell D. Ebbing, David Bookin**



## **Chemistry Equilibrium Practice Problems With Answers:**

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