



CHEMISTRY

CHEAT SHEET

GAS LAWS



GAS LAWS CHEAT SHEET

Remember: a gas is a substance composed of particles which move from one volume to another (diffuse) and exert a pressure on the container. Boyle's Law, Charles' Law, and Gay-Lussac's Law are the three gas laws that describe the behavior of a gas.

Boyle's Law
Charles' Law
Gay-Lussac's Law

Boyle's Law	Charles' Law	Gay-Lussac's Law
Pressure (P)	Volume (V)	Temperature (T)
Volume (V)	Pressure (P)	Volume (V)
Temperature (T)	Temperature (T)	Pressure (P)
Pressure (P)	Volume (V)	Temperature (T)

Boyle's Law

Remember: pressure and volume are inversely related.

$$P_1 V_1 = P_2 V_2$$

Remember: as pressure increases, volume decreases, and vice versa.



Charles' Law

Remember: volume and temperature are directly related.

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

Remember: as temperature increases, volume increases, and vice versa.



Gay-Lussac's Law

Remember: pressure and temperature are directly related.

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

Remember: as temperature increases, pressure increases, and vice versa.



Ideal Gas Law

The ideal gas law is a combination of Boyle's, Charles', and Gay-Lussac's Laws. It is written as follows:

$$PV = nRT$$

P = pressure (atm) V = volume (L) n = number of moles R = gas constant (0.0821 L·atm/mol·K)

Remember: the units must cancel.

$$P = \frac{nRT}{V}$$

$$V = \frac{nRT}{P}$$

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