

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium ($\text{H}-3$) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{aligned} 63 \times .692 &= 43.596 \rightarrow 44 \\ 65 \times .308 &= 20.02 \rightarrow 20 \\ + & \\ \hline 63.62 &\rightarrow 64 \end{aligned}$$

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Susan McMurry



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Table of Contents Chemistry Guide Answers

1. Understanding the eBook Chemistry Guide Answers
 - The Rise of Digital Reading Chemistry Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Guide Answers
 - Personalized Recommendations
 - Chemistry Guide Answers User Reviews and Ratings
 - Chemistry Guide Answers and Bestseller Lists
5. Accessing Chemistry Guide Answers Free and Paid eBooks
 - Chemistry Guide Answers Public Domain eBooks
 - Chemistry Guide Answers eBook Subscription Services
 - Chemistry Guide Answers Budget-Friendly Options

6. Navigating Chemistry Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Guide Answers Compatibility with Devices
 - Chemistry Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Guide Answers
 - Highlighting and Note-Taking Chemistry Guide Answers
 - Interactive Elements Chemistry Guide Answers
8. Staying Engaged with Chemistry Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Guide Answers
9. Balancing eBooks and Physical Books Chemistry Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Chemistry Guide Answers
 - Carving Out Dedicated Reading Time
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 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
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