- 25. The number of  $\sigma$  bonds in N $\equiv$ N is
  - die. 2 Mr.
  - 34
  - 4 d.
- 26. The elements in an ionic compound are held together by
  - electrostatic forces of attraction. ar.
  - van der Waals forces
  - the spin of paired electrons.
  - the formation of hybrid orbitals. di
  - an electron pair.
- 27. In every electrolytic and galvanic (voltaic) cell the anode is that electrode
  - at which oxidation occurs.
  - which attracts cations.
  - at which electrons are supplied to the
  - at which reduction occurs.
- 28. Metal X was plated from a solution containing cations of X. The passage of 48.25 C deposited 31 mg of X on the cathode. What is the mass of X (in grams) per mole of electrons?
  - ec: 4.7
  - Sec. 62:
  - 903
  - 1865 de
- In a galvanic (voltaic) cell in which the  $Cd + Cu^{2+} \rightarrow Cu + Cd^{2+}$ reaction is and the ions are at unit concentration (activity), the cell potential is

$$Cd \rightarrow Cd^{2*} + 2e$$
 0.4021 V  
 $Cu \rightarrow Cu^{2*} + 2e$  0.344 V

- a. 0.1383 V
- A. 0.4021 V. e. 0.344 V
- d. 0.7461 V
- 0.3677 V
- 340. In which reaction will an increase in total pressure at constant temperature favor formation of the products?
  - a. CaCO₂(s) ←→ CaO(s) + CO₂(g)
  - b. H<sub>2</sub>(g) + Cl<sub>2</sub>(g) ← → 2HCl(g)
  - c.  $2NO(g) + O_2(g) \leftarrow \rightarrow 2NO_2(g)$
  - d. COCl₂(g) ← → CO(g) + Cl₂(g)

Standard Potentials	E
$Mg \rightarrow Mg^{2} + 2e$	2.37V
Al → Al <sup>5</sup> *+3e	1.66V
$Zn \rightarrow Zn^{2*}+2e$	0.76V
Fe⇒Fe <sup>1a</sup> +2e	0.44V
$Cu \rightarrow Cu^{2s} + 2e$	0.34V
Ag→ Ag+e	-0.80V

- 31. Using only the metals Mg, Al, Zn, Fe, Cu and Ag, together with their 1 M salt solutions, a voltaic cell of the highest possible voltage would be constructed using electrodes of these metals.
  - Mg and Ag.
  - Mig and Fe
  - Zn and Cu
  - Al and Ag
  - Mg and Al
- 32.  $E = E^{\circ} - 0.059/n \log Q$  (Nernst equation)  $[H^*] = 1.0 \text{ M initially}, P_{e2} = 1.0 \text{ atm}$

 $4c + O_2(g) + 4H'(aq) \leftarrow \rightarrow 2H_2O(1)$ E'=1.23VBased on the information above, which statement is correct?

- n = 1, since one mole of oxygen is being considered.
- Addition of base should result in an E value, which is less than 1.23 V.
- E is independent of the pH of the solution.
- $Q = [H_2Q]^2$ d. IO:HH!
- 33. The equilibrium constant for the gaseous reaction C + D ← → E + 2F is 3.0 at 50 °C. In a 2.0 L flask at 50 °C are placed 1.0 mol of C, 1.0 mol of D, 1.0 mol of E, and 3.0 mol of F. Initially, the reaction will
  - proceed at equal rates in both directions.
  - proceed more rapidly to form E and F.
  - proceed more rapidly to form C and D.
  - not occur in either direction.

Compound	ΔG° <sub>ε</sub> kJ/mol	
H <sub>2</sub> O(1)	-237	
H <sub>2</sub> O(g)	- 229	

- At 298 K the equilibrium constant for 34.  $H_2(g) + 55 O_2(g) \leftarrow \rightarrow H_2O(1)$ 
  - is larger than the Keq for  $H_2(g) + \frac{1}{2} O_2(g) \leftarrow \rightarrow H_2O(g)$
  - will have a value of 10 at equilibrium.
  - cannot be computed since data on O2 and

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**CL Gary** 

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