

Conversion factors:

Molar mass = #g/1 mole

Avogadro's number = 6.02×10^{23} atoms or molecules/1 mole

Molarity = #moles/L

Name _____

Period _____

Stoichiometry Practice Problem Set

1. In the following reaction $4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$, iron oxidizes to form rust.
 - a) If 40.0g of iron react, how many moles of Fe_2O_3 , iron (III) oxide are produced?
 - b) If 20.0g of iron react, how many grams of Fe_2O_3 are produced?
 - c) If 10.0 moles of iron react, how many molecules of Fe_2O_3 are produced?
 - d) If 3.00×10^{22} atoms of iron react, how many moles of Fe_2O_3 are produced?
2. In the reaction for photosynthesis, $6\text{CO}_2(g) + 6\text{H}_2\text{O}(l) + \text{light energy} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(aq) + 6\text{O}_2(g)$,
 - a) If 715g of CO_2 gas is consumed, how many grams of glucose ($\text{C}_6\text{H}_{12}\text{O}_6$) are produced? The molar mass of glucose is 180.16g/mol.
 - b) How many molecules of carbon dioxide are required to produce 5.9×10^{24} molecules of oxygen gas?

Chemistry Stoichiometry Problem Set 3 Answers

Chris Sunley



Chemistry Stoichiometry Problem Set 3 Answers:

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