

4. a.
$$\frac{10 \text{ g Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 35.33 \text{ g Cu}$$

$$\frac{20 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol Cu}} = \boxed{9.45 \text{ g Cu}}$$

b.
$$\frac{30 \text{ g Al} \mid 1 \text{ mol Al} \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol AlCl}_3} = 148.25 \text{ g AlCl}_3$$

$$\frac{100 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = \boxed{66.11 \text{ g AlCl}_3}$$

c.
$$\frac{5 \text{ g Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 17.67 \text{ g Cu}$$

~~100 g / 17.67 g~~ $10 \text{ g} / 17.67 \text{ g} \times 100 = \boxed{56.6\%}$

d.
$$\frac{10 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = 6.61 \text{ g AlCl}_3$$

~~100 g / 6.61 g~~ $5 \text{ g} / 6.61 \text{ g} \times 100 = \boxed{75.6\%}$

- 5.
- a. TRUE
 - b. FALSE
 - c. $\text{AgNO}_3 \Rightarrow 169.91 \text{ g/mol}$
 - d. 6.02×10^{23} particles = 1 mole of particles

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Donald J. DeCoste



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