# Mole Concept

1 mole of a thing = 6.02214 x 10<sup>23</sup> of those things



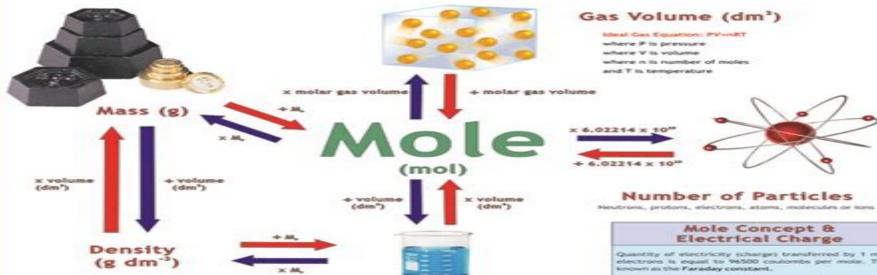
To carry out a reaction that requires combining one atom of iron with one atom of sulfur, how much iron should be used? How much sulfur? In the lab, there is no device that can count numbers of atoms. Besides, the merest speck (0.001 g) of iron contains over a billion billion atoms. The same is true of sulfur.

The mole is a chemist's measure of the amount of substance.

One mole is defined as the amount of substance containing the same number of chemical units (atoms, molecules or ions) as there are atoms in exactly 12 grams of pure carbon 12.

The number of atoms present in 12 grams of carbon 12 is equal to 6.02214 X 10<sup>10</sup>.

This number is also known as Avogadro's constant (N.).



#### Solids Counting the Atoms by Weighing them

Ovygon Magnesium Marcs Combined 266 Manhair Manna 2-lig/ym 56g/mol Number of moi ti smoit. Soom P. Simplest ratio Formula. MAUO



When 24 grams of magnesium is burnt. then 40 grams of Magnesium Oxide is formed. By Law of conservation of mass, we conclude 16 grams of oxygen is used. Hence the formula MgO.

A chemical equation indicates the number of moles of reactants and products involved in the reaction. This helps in determining the quantity of reactants required to manufacture desired quantity of a e Supremier at

### (mol dm<sup>-3</sup>) Liquids

Concentration

When a chemical substance or the solute is dissolved in a volume of solvent, the mass concentration of a solution is measured in grams per cubic declimetre (g/dm/). The concentration of a solution is measured in moles per cubic decimetre (mot/dm'). When I mot of a substance is dissolved in 1 dm' of water, a solution with a concentration of 1 moil/dm' is produced.

Number of moles in a solution = Concentration x volume of solution (in cm')

where concentration is in moles per cubic decimetre. but volume of solution is in cubic centimetres.

1 mol of copper sulphase CuSO.

#### Mole Concept & **Electrical Charge**

Quantity of electricity (charge) transferred by 1 mol of electrons is equal to 96500 coulombs per mole. This is

Charge on one mole of electrons = 96500 coulombs.

The Faraday constant = 96500 C/mol.

#### Gases Avogadro's Law

Avogadro's Law: 'Under the similar conditions of temperature and pressure, equal volumes of all pages contain same number of molecules".

This number (Avogadro's number) is 6.02214 X 10". It is the number of molecules of an ideal gas present in a volume of 22.41 dm" at 5.T.P. (standard temperature &

Reacting Hydrogen gas with Chlorine gas

Hydrogen + Chlorine ---94, 0gb CL tg1 2 HICH END -22.41dm -22.41dm =44.82dm



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**George Gilbert Long, Forrest C. Hentz** 

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