

Mole Concept

1 mole of a thing = 6.02214×10^{23} of those things



Amedeo Avogadro

To carry out a reaction that requires combining one atom of iron with one atom of sulfur, how much iron should be used? How much sulfur? In the lab, there is no device that can count numbers of atoms. Besides, the merest speck (0.001 g) of iron contains over a billion billion atoms. The same is true of sulfur.

The mole is a chemist's measure of the amount of substance.

One mole is defined as the amount of substance containing the same number of chemical units (atoms, molecules or ions) as there are atoms in exactly 12 grams of pure carbon 12.

The number of atoms present in 12 grams of carbon 12 is equal to 6.02214×10^{23} .

This number is also known as **Avogadro's constant** (N_A).



Gas Volume (dm³)

Ideal Gas Equation: $PV=nRT$

where P is pressure
where V is volume
where n is number of moles
and T is temperature

$$\begin{aligned} &\times 6.02214 \times 10^{23} \\ &\div 6.02214 \times 10^{23} \end{aligned}$$



Number of Particles

Neutrons, protons, electrons, atoms, molecules or ions

Mole Concept & Electrical Charge

Quantity of electricity (charge) transferred by 1 mol of electrons is equal to 96500 coulombs per mole. This is known as the **Faraday constant**.

Charge on one mole of electrons = 96500 coulombs.

The Faraday constant = 96500 C/mol.

Solids

Counting the Atoms by Weighing them

	Magnesium	Oxygen
Mass Combined	24g	16g
Molar Mass	24g/mol	16g/mol
Number of moles	1 mol	1 mol
Simpliest ratio	1	1
Formula	MgO	



When 24 grams of magnesium is burnt, then 40 grams of Magnesium Oxide is formed. By law of conservation of mass, we conclude 16 grams of oxygen is used. Hence the formula MgO.

A chemical equation indicates the number of moles of reactants and products involved in the reaction. This helps in determining the quantity of reactants required to manufacture desired quantity of a chemical.

Liquids

When a chemical substance or the solute is dissolved in a volume of solvent, the mass concentration of a solution is measured in grams per cubic decimetre (g/dm³). The concentration of a solution is measured in moles per cubic decimetre (mol/dm³). When 1 mol of a substance is dissolved in 1 dm³ of water, a solution with a concentration of 1 mol/dm³ is produced.

Number of moles in a solution = $\frac{\text{Concentration} \times \text{volume of solution (in cm}^3\text{)}}{1000}$

where concentration is in moles per cubic decimetre, but volume of solution is in cubic centimetres.

1 mol of copper sulphate CuSO₄



Gases

Avogadro's Law

Avogadro's Law: "Under the similar conditions of temperature and pressure, equal volumes of all gases contain same number of molecules".

This number (Avogadro's number) is 6.02214×10^{23} . It is the number of molecules of an ideal gas present in a volume of 22.41 dm³ at S.T.P. (standard temperature & pressure).

Reacting Hydrogen gas with Chlorine gas



Chemistry Answers The Mole Concept

George Gilbert Long, Forrest C. Hentz



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