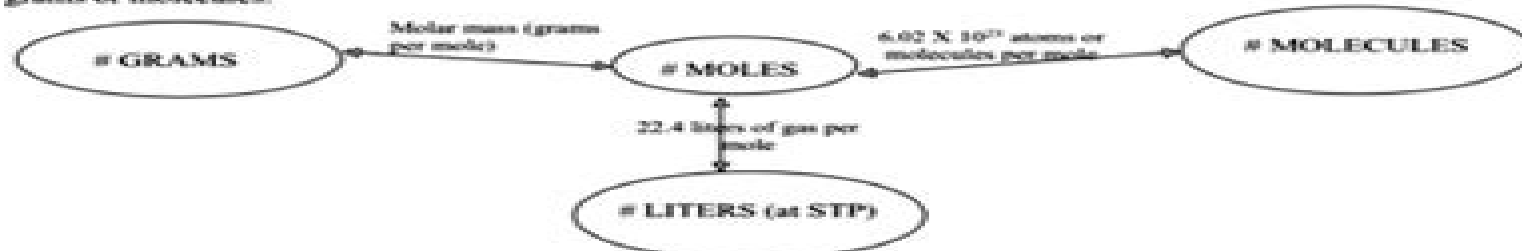


You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP. We can use this information to convert grams to molecules or liters, molecules to grams or liters, or liters to grams or molecules.



EXAMPLE 1: What would be the volume in liters of 40.36 grams of neon at STP?

$$\# \text{ liters Ne} = 40.36 \text{ g Ne} \times \frac{1 \text{ mole Ne}}{20.18 \text{ g Ne}} \times \frac{22.4 \text{ L Ne}}{1 \text{ mole Ne}} = 44.80 \text{ L Ne}$$

EXAMPLE 2: How many molecules would there be in 56 liters of carbon dioxide at STP?

$$\# \text{ molecules CO}_2 = 56.0 \text{ L CO}_2 \times \frac{1 \text{ mole CO}_2}{22.4 \text{ L CO}_2} \times \frac{6.02 \times 10^{23} \text{ molecules CO}_2}{1 \text{ mole CO}_2} = 1.51 \times 10^{24} \text{ molecules CO}_2$$

SOLVE THE FOLLOWING PROBLEMS ON A SEPARATE SHEET OF PAPER.

- YOU MUST USE COMPLETE AND PROPER SET-UPS.
- SHOW THE MOLAR MASS CALCULATION WHENEVER THE PROBLEM REQUIRES YOU TO DO ONE. INCLUDE UNITS.
- CIRCLE YOUR FINAL ANSWER.

1. What would be the volume, in liters, of 85.5 grams of carbon monoxide at STP?
2. How many molecules would there be in 0.500 grams of carbon disulfide?
3. What would be the mass, in grams, of 45.0 liters of nitrogen at STP?
4. How many molecules of hydrogen are in a balloon full of hydrogen with a volume of 5.34 liters at STP?
5. Your mommy buys you a helium balloon at the circus. It has a volume of 4.00 liters at STP. What mass of helium, expressed in grams, does this balloon contain?
6. How many molecules of ammonia would there be in 40.0 grams of ammonia?
7. What would be the mass, in grams, of 3.50×10^{23} molecules of chlorine?
8. What volume, expressed in liters, would 50.0 grams of fluorine occupy at STP?
9. How many grams of oxygen would there be in 1.00 liter of oxygen at STP?

Extra Challenge:

10. How many molecules of water are there in 10 lbs of water?

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Table of Contents Chemistry B Moles Packet Worksheet 6

1. Understanding the eBook Chemistry B Moles Packet Worksheet 6
 - The Rise of Digital Reading Chemistry B Moles Packet Worksheet 6
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry B Moles Packet Worksheet 6
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry B Moles Packet Worksheet 6
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry B Moles Packet Worksheet 6
 - Personalized Recommendations
 - Chemistry B Moles Packet Worksheet 6 User Reviews and Ratings
 - Chemistry B Moles Packet Worksheet 6 and Bestseller Lists
5. Accessing Chemistry B Moles Packet Worksheet 6 Free and Paid eBooks
 - Chemistry B Moles Packet Worksheet 6 Public Domain eBooks
 - Chemistry B Moles Packet Worksheet 6 eBook Subscription Services
 - Chemistry B Moles Packet Worksheet 6 Budget-Friendly Options
6. Navigating Chemistry B Moles Packet Worksheet 6 eBook Formats

- ePub, PDF, MOBI, and More
 - Chemistry B Moles Packet Worksheet 6 Compatibility with Devices
 - Chemistry B Moles Packet Worksheet 6 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry B Moles Packet Worksheet 6
 - Highlighting and Note-Taking Chemistry B Moles Packet Worksheet 6
 - Interactive Elements Chemistry B Moles Packet Worksheet 6
 8. Staying Engaged with Chemistry B Moles Packet Worksheet 6
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry B Moles Packet Worksheet 6
 9. Balancing eBooks and Physical Books Chemistry B Moles Packet Worksheet 6
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry B Moles Packet Worksheet 6
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry B Moles Packet Worksheet 6
 - Setting Reading Goals Chemistry B Moles Packet Worksheet 6
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry B Moles Packet Worksheet 6
 - Fact-Checking eBook Content of Chemistry B Moles Packet Worksheet 6
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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