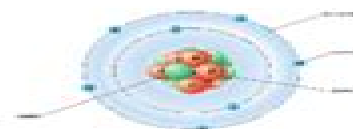


GCSE Chemistry AQA C1 Atomic Structure & Periodic Table

Atoms are **neutral** particles which make up elements

Nuclear Model of Atom is small central **nucleus** of positively charged **protons** (relative charge +1) and neutral **neutrons** (relative charge 0) surrounded by very small negatively charged **electrons** (relative charge -1) orbiting in shells



Ion is **charged particle** where an atom has lost or gained electrons (to obtain a full outer shell)
 e.g. sodium ion (Na^+) which was sodium but lost 1 negative electron so is now more positive
 e.g. magnesium (Mg^{2+}) which was magnesium but lost 2 negative electrons so is now more positive
 e.g. oxide ion (O^{2-}) which was oxygen but gained 2 negative electrons so is now more negative

Elements are substances containing **only one type of atom**, e.g. sodium: (Na), iodine (I_2), oxygen (O_2)

Compounds are substances containing **two or more atoms chemically combined together in fixed proportions**, e.g. sodium chloride (NaCl)

Mixtures are substances containing **two or more atoms not chemically combined together** with chemical properties unchanged, e.g. alloys (metals and other solids), rock salt (sand and salt), air

Subatomic Particles: protons and neutrons (both relative mass 1), electrons (relative mass very small)

Mass Number is number of protons and neutrons in nucleus of an atom



NOTE: Iron is $^{56}_{26}\text{Fe}$ = 4 times heavier than nitrogen

Atomic Number is number of **protons** (= same number of **electrons**) in an atom

Number of Neutrons is difference between the mass number and atomic number

Boyle suggested some substances cannot be made simpler; these are **elements**

Dalton suggested each element was composed of indivisible **spherical atoms** of same mass

JJ Thomson suggested atoms were **negative electrons embedded in positive spheres of charge** (plum pudding model)

Boyle/Dalton model



Plum pudding model



Rutherford's model



Bohr's / Modern Nuclear model



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DP Hallahan



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