

# Chemistry

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(Chapter 2)(Solutions)  
(Class 12)  
Intext Questions

## Question 2.1:

Calculate the mass percentage of benzene ( $C_6H_6$ ) and carbon tetrachloride ( $CCl_4$ ) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

## Answer 2.1:

$$\text{Mass percentage of } C_6H_6 = \frac{\text{Mass of } C_6H_6}{\text{Total mass of the solution}} \times 100\%$$

$$= \frac{\text{Mass of } C_6H_6}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100\% = \frac{22}{22 + 122} \times 100\% = 15.28\%$$

$$\text{Mass percentage of } CCl_4 = \frac{\text{Mass of } CCl_4}{\text{Total mass of the solution}} \times 100\%$$

$$= \frac{\text{Mass of } CCl_4}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100\% = \frac{122}{22 + 122} \times 100\% = 84.72\%$$

Alternatively, Mass percentage of  $CCl_4 = (100 - 15.28)\% = 84.72\%$

## Question 2.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

## Answer 2.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

$\therefore$  Mass of carbon tetrachloride =  $(100 - 30)$  g = 70 g

Molar mass of benzene ( $C_6H_6$ ) =  $(6 \times 12 + 6 \times 1)$  g mol<sup>-1</sup> = 78 g mol<sup>-1</sup>

$\therefore$  Number of moles of  $C_6H_6 = \frac{30}{78}$  mol = 0.3846 mol

Molar mass of carbon tetrachloride ( $CCl_4$ ) =  $1 \times 12 + 4 \times 35.5 = 154$  g mol<sup>-1</sup>

$\therefore$  Number of moles of  $CCl_4 = \frac{70}{154}$  mol = 0.4545 mol

Thus, the mole fraction of  $C_6H_6$  is given as:

$$\frac{\text{Number of moles of } C_6H_6}{\text{Number of moles of } C_6H_6 + \text{Number of moles of } CCl_4} = \frac{0.3846}{0.3846 + 0.4545} = 0.458$$

## Question 2.3:

Calculate the molarity of each of the following solutions: (a) 30 g of  $Co(NO_3)_2 \cdot 6H_2O$  in 4.3 L of solution (b) 30 mL of 0.5 M  $H_2SO_4$  diluted to 500 mL.

## Answer 2.3:

Molarity is given by:  $\text{Molarity} = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$

(a) Molar mass of  $Co(NO_3)_2 \cdot 6H_2O = 59 + 2(14 + 3 \times 16) + 6 \times 18 = 291$  g mol<sup>-1</sup>

$\therefore$  Moles of  $Co(NO_3)_2 \cdot 6H_2O = \frac{30}{291}$  mol = 0.103 mol

Therefore, molarity =  $\frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$

(b) Number of moles present in 1000 mL of 0.5 M  $H_2SO_4 = 0.5$  mol

$\therefore$  Number of moles present in 30 mL of 0.5 M  $H_2SO_4 = \frac{0.5 \times 30}{1000}$  mol = 0.015 mol

Therefore, molarity =  $\frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$

## Question 2.4:

Calculate the mass of urea ( $NH_2CONH_2$ ) required in making 2.5 kg of 0.25 molar aqueous solution.

## Answer 2.4:

Molar mass of urea ( $NH_2CONH_2$ ) =  $2(1 \times 14 + 2 \times 1) + 1 \times 12 + 1 \times 16 = 60$  g mol<sup>-1</sup>

0.25 molar aqueous solution of urea means: 1000 g of water contains 0.25 mol =  $(0.25 \times 60)$  g of urea = 15 g of urea

That is,

(1000 + 15) g of solution contains 15 g of urea

Therefore, 2.5 kg (2500 g) of solution contains =  $\frac{15 \times 2500}{1000 + 15}$  g = 36.95 g = 37 g of urea (approximately)

Hence, mass of urea required = 37 g

# Chemistry Chapter 2 Review Measurements And Calculations Answer Key

**David E. Goldberg**



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