

Name: _____

AP Chemistry Chemical Equations Worksheet

Write the balanced chemical equation (excluding spectator ions) underneath each reaction description, and answer the question.

(a) A solution of ammonia is added to a dilute solution of acetic acid.	Identify the conjugate acid-base pairs in this reaction.
(b) Solutions of sodium hydroxide and acetic acid are mixed.	Acetic acid is a weak acid. If equal volumes of equal molar solutions are mixed, will the solution pH be >7, 7, or <7? Explain.
(c) Hydrogen sulfide gas is bubbled through excess potassium hydroxide solution.	Write the successive ionization equations for H_2S .
(d) Solid barium oxide is added to distilled water.	Is the resulting solution acidic, basic or neutral? Explain.
(e) Solid calcium oxide is exposed to a stream of carbon dioxide gas.	What type of reaction has occurred?
(f) Solid dinitrogen pentoxide is added to water.	Is the final solution acidic, basic or neutral? Explain.
(g) Carbon disulfide vapor is burned in excess oxygen.	What is the oxidizing agent in this reaction?

Chemistry Chemical Equations Flinn Answer Key

Sebastian Brünink



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CHEMISTRY SILBERBERG,2003 *AP Chemistry* Albert S. Tarendash,Frederick J. Rowe,2002 This indispensable guide to chemistry helps students who wish to prepare for the AP Chemistry exam on their own Comprehensive and easy to understand this learning guide includes a full content review two full length practice tests with hundreds of practice questions and thorough answer explanations and proven test taking strategies **Corrosion and Electrochemistry of Zinc** Xiaoge Gregory Zhang,2013-06-29 Humankind s use of zinc stretches back to antiquity and it was a component in some of the earliest known alloy systems Even though metallic zinc was not discovered in Europe until 1746 by Marggral zinc ores were used for making brass in biblical times and an 87% zinc alloy was found in prehistoric ruins in Transylvania Also zinc the metal was produced in quantity in India as far back as the thirteenth century well before it was recognized as being a separate element The uses of zinc are manifold ranging from galvanizing to die castings to electronics It is a preferred anode material in high energy density batteries e g Ni Zn Ag Zn ZnJair so that its electrochemistry particularly in alkaline media has been extensively explored In the passive state zinc is photoelectrochemically active with the passive film displaying n type characteristics For the same reason that zinc is considered to be an excellent battery anode it has found extensive use as a sacrificial anode for the protection of ships and pipelines from corrosion Indeed aside from zinc s well known attributes as an alloying element its widespread use is principally due to its electrochemical properties which include a well placed position in the galvanic series for protecting iron and steel in natural aqueous environments and its reversible dissolution behavior in alkaline solutions **Chemistry in the Laboratory** James M. Postma,Julian L. Robert,J. Leland Hollenberg,2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

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Chemistry John S. Phillips,Cheryl Wistrom,2000 *The Hotel World* ,1919 [Children's Books in Print](#) ,2007 ,2006

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