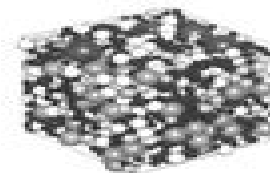




Key

Density Practice Problems



The density of a substance is a measure of how much mass is packed into a certain volume of the substance. Substances with a high density, like steel, have molecules that are packed together tightly. Substances with a low density, like cork, have fewer molecules packed into the same amount of space.

The density of a substance can be found by dividing its mass by its volume. As long as a substance is homogeneous, the size or shape of the sample doesn't matter. The density will always be the same. This means that a steel paper clip has the same density as a steel girder used to build a bridge.

$$\text{Density} = \text{Mass} \div \text{Volume}$$

$$D = \frac{m}{V}$$

Use the density formula to solve the following problems. Show all work and the answer must have the correct units. Remember that volume can have different forms. A block of ice with a volume of 3 cm³ would be 3 mL of liquid after being melted.

1. What is the density of CO gas if 0.196 g occupies a volume of 100 ml?

$$m = 0.196 \text{ g}$$

$$V = 100 \text{ ml}$$

$$D = \frac{0.196 \text{ g}}{100 \text{ ml}}$$

$$\text{Answer: } 0.00196 \text{ g/ml} = 0.002 \text{ g/ml}$$

2. A block of wood 3 cm on each side has a mass of 27 g. What is the density of the block? (Hint, don't forget to find the volume of the wood.)

$$V = L \times W \times H$$

$$V = 3 \times 3 \times 3$$

$$V = 27 \text{ cm}^3$$

$$D = \frac{27 \text{ g}}{27 \text{ cm}^3}$$

$$D = 1 \text{ g/cm}^3$$

$$m = 27 \text{ g}$$

$$\text{Answer: } 1 \text{ g/cm}^3$$

3. A student measures a rectangular block of an unknown substance. The block measures 12.1 cm by 17.5 cm by 5.8 cm. The student weighs the block and determines its mass to be 4250.112 g. What is the density of this material? Watch the sig figs.

$$V = L \times W \times H$$

$$V = 12.1 \times 17.5 \times (5.8) \quad 2 \text{ s.f.}$$

$$V = 1228.15 \text{ cm}^3$$

$$D = \frac{4250.112 \text{ g}}{1228.15 \text{ cm}^3}$$

$$\text{Answer: } 3.46 \text{ g/cm}^3 = 3.5 \text{ g/cm}^3$$

4. An irregularly shaped stone was lowered into a graduated cylinder holding a volume of water equal to 3.5 ml. The height of the water rose to 8.0 ml. If the mass of the stone was 27 g, what was its density?

$$V = 8.0 \text{ mL} - 3.5 \text{ mL}$$

$$m = 27 \text{ g}$$

$$V = 4.5 \text{ mL}$$

$$D = \frac{27 \text{ g}}{4.5 \text{ mL}} = 6$$

$$\text{Answer: } 6.0 \text{ g/mL}$$

Chemistry Density Problems With Answers

VM Jensen



Chemistry Density Problems With Answers:

Dr. Salm's Chemistry Problem Drill Book Robert J. Salm, Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding

Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

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