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Given that

$$H_2(g) + F_2(g) \longrightarrow 2HF(g)$$

$$\Delta H_{cm}^* = -546.6 \, \text{kJ·mol}^{-1}$$

$$2H_2(g) + O_2(g) \longrightarrow 2H_2O(f)$$

$$\Delta H_{\rm res}^2 = -571.6 \, \text{kJ·mol}^{-1}$$

calculate the value of ΔH_{cos} for

$$2F_2(g) + 2H_2O(l) \longrightarrow 4HF(g) + O_2(g)$$

$$\Delta H_{\rm rin}^* =$$

ki-mol*1

Question Source McQuartie, Buck, And Gallegiy Ser-Control Chemistry | Publisher University Science Books

































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John Ashhurst

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