Electronegativity

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Information: Definition of Electronegativity

Electronogativity is a recasure of how much an atom attracts an electron. The higher the electronogativity, the greater the atom's attraction for electrons. Atoms that become negative ions have a reach greater electronogativity than atoms that become positive ions. Below is a table of the electronogativities of many elements from the periodic table.

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Critical Thinking Questions

- Why do you think there are no values for the noble gases?
 Noble gases are so stable that they do not attract electrons.
- In terms of electrons, what is the difference between a covalent bond and an ionic bond?
 Electrons are shared in a covalent bond, but arona loss or sain there is loss to bonds.
- What type of bond (covalent or ionic) would you expect to form between an atom with a high electronogativity and an atom of low electronogativity. Explain and give as example.
 Ionic bond because the atom with high electronogativity will attract an electron away from the atom with low electronogativity. Example: NaF Outween a rectal and a summetal)
- 4. What type of bond (covalent or ionic) would you expect to form between two atoms of somewhat high electronegativity? Explain and give an example.
 Covalent bond because both atoms will be trying to attract an electron and so they will end apartning it. Example: CO₂ (both atoms have high electronegativity compared to the metals)
- 5. Consider the ionic compound, welliam chloride (NaCl). Which atom has a greater attraction for electrons—sodium or oblorine? Which atom forms the negative ion? Chlorine has the highest electronegativity and therefore the strongest attraction. It also forms the negative ion and addison forms a positive ion.

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