### Experiment 13

# Advance Study Assignment: The Geometrical Structure of Molecules

You are asked by your instructor to construct a model of the CH<sub>3</sub>Cl molecule. Being of a conservative nature, you proceed as directed in the section on Experimental Procedure.

 First you need to find the number of valence electrons in CH<sub>3</sub>Cl. The number of valence electrons in an atom of an element is equal to the last digit in the group number of that element in the Periodic Table.

C is in Group \_\_\_\_\_ Cl is in Group \_\_\_\_\_

In CH,Cl there is a total of \_\_\_\_\_\_ valence electrons.

- 2. The model consists of balls and sticks. What kind of ball should you select for the C atom? \_\_\_\_\_ the H atoms? \_\_\_\_ the Cl atom? \_\_\_\_ The electrons in the molecule are paired, and each stick represents an electron pair. How many sticks do you need? \_\_\_\_
- Assemble a skeleton structure for the molecule, connecting the balls with sticks into one unit. Use the
  rule that C atoms form four bonds, whereas Cl atoms usually do not. Draw a sketch of the skeleton
  below:
- 4. How many sticks did you need to make the skeleton structure? \_\_\_\_\_\_ How many sticks are left over? \_\_\_\_\_ If your model is to obey the octet rule each ball must have four sticks in it (except for hydrogen atom balls, which need only one). (Each atom in an octet rule species is surrounded by four pairs of electrons.) How many holes remain to be filled? \_\_\_\_\_ Fill them with the remaining sticks, which represent nonbonding electron pairs. Draw the complete Lewis structure for CH3Cl using lines for bonds and pairs of dots for nonbonding electrons.
- Describe the geometry of the model, which is that of CH<sub>3</sub>Cl. \_\_\_\_\_\_ Is the CH<sub>3</sub>Cl molecule polar? \_\_\_\_\_\_ Why?

Would you expect CH3Cl to have any isomeric forms? \_\_\_\_\_ Explain your reasoning.

6. Would CH<sub>3</sub>Cl have any resonance structures? \_\_\_\_\_ If so, draw them below.

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**R Bogdan** 

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