

2

MATTER AND CHANGE

SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)

This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.

► Describing Matter (page 39)

1. The _____ mass _____ of an object is a measure of the amount of matter the object contains.

2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

► Identifying Substances (page 40)

3. Matter that has a uniform and definite composition is called a _____ substance _____.

4. Is the following sentence true or false? All samples of a substance have different physical properties. _____ false _____

5. A physical property is a quality or condition of a substance that can be _____ observed _____ or _____ measured _____ without changing the substance's composition.

6. Circle the letter of the term that is NOT a physical property.

- a. hardness c. boiling point
b. color **d. melting**

7. Look at Table 2.1 on page 40. What is the melting point of bromine? _____ -7°C _____

8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at 115°C.

- a. sulfur**
b. chlorine
c. gold
d. copper

Chemistry Matter And Change Chapter 1study Guide Answers

Susan McMurry



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1. Understanding the eBook Chemistry Matter And Change Chapter 1study Guide Answers
 - The Rise of Digital Reading Chemistry Matter And Change Chapter 1study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Chapter 1study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Matter And Change Chapter 1study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Chapter 1study Guide Answers
 - Personalized Recommendations
 - Chemistry Matter And Change Chapter 1study Guide Answers User Reviews and Ratings

- Chemistry Matter And Change Chapter 1study Guide Answers and Bestseller Lists
- 5. Accessing Chemistry Matter And Change Chapter 1study Guide Answers Free and Paid eBooks
 - Chemistry Matter And Change Chapter 1study Guide Answers Public Domain eBooks
 - Chemistry Matter And Change Chapter 1study Guide Answers eBook Subscription Services
 - Chemistry Matter And Change Chapter 1study Guide Answers Budget-Friendly Options
- 6. Navigating Chemistry Matter And Change Chapter 1study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Matter And Change Chapter 1study Guide Answers Compatibility with Devices
 - Chemistry Matter And Change Chapter 1study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Matter And Change Chapter 1study Guide Answers
 - Highlighting and Note-Taking Chemistry Matter And Change Chapter 1study Guide Answers
 - Interactive Elements Chemistry Matter And Change Chapter 1study Guide Answers
- 8. Staying Engaged with Chemistry Matter And Change Chapter 1study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Matter And Change Chapter 1study Guide Answers
- 9. Balancing eBooks and Physical Books Chemistry Matter And Change Chapter 1study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter And Change Chapter 1study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Matter And Change Chapter 1study Guide Answers
 - Setting Reading Goals Chemistry Matter And Change Chapter 1study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Matter And Change Chapter 1study Guide Answers
 - Fact-Checking eBook Content of Chemistry Matter And Change Chapter 1study Guide Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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