

- (B) decrease by 2 times
- (C) increase by 2 times
- (D) remain the same

Answer (A)

Sol. Rate = $k [A]^1[B]^2$
 It is 1st order w.r.t A and 2nd order w.r.t. B
 If B is doubled so rate will increase by = $2^2 = 4$ times

11. Photochemical reactions like formation of HCl from H₂ and Cl₂ are ____.

- (A) zero order reactions
- (B) first order reactions
- (C) second order reactions
- (D) third order reactions

Answer (A)

Sol. $H_2 + Cl_2 \rightarrow HCl$
 All photochemical reactions, some enzyme catalyzed reactions and reactions which occur on metal surfaces are zero order reactions.

12. A first order reaction is half completed in 3 minutes. How long does it take for completion of 99% of the reaction? (Given: $2.303/0.231 = 9.969$; $\log 100 = 2.0000$)

- (A) 1.9938 minutes
- (B) 19.938 minutes
- (C) 9.969 minutes
- (D) 91.93 minutes

Answer (B)

Sol. $k = 2.303/t \times \log(A_0/A)$
 A_0 = Initial concentration
 A = Final concentration
 $k = 0.693/t_{1/2} = 0.693/3$
 $t = 2.303/k \times \log(A_0/A) = 2.303 \times 3/0.693 \log(100/1)$
 $t = 19.938$ minutes

13. If the rate constant of a reaction is $0.3 \times 10^{-4} \text{ mol}^{-1} \text{ L s}^{-1}$, the order of the reaction is ____.

- (A) zero order
- (B) first order
- (C) second order
- (D) third order

Answer (C)

Sol. Rate of reaction = $k [\text{concentration}]^n$
 $k = \text{Rate of reaction}/[\text{concentration}]^n = \text{mol L}^{-1} \text{ s}^{-1}/(\text{mol L}^{-1})^2$
 Unit of $k = \text{mol}^{-1} \text{ L s}^{-1}$ indicates that it is a second order reaction.

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Mrs. Helen Cecilia De Silver (Abbott)
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