

Common Error

-  Molar volume is used to determine the volume of a solution.
-  Molar volume is used to determine the volume of a gas. The volume of a solution can be determined using its concentration and the number of moles of solute present in the solution.

 **Link** → Chemistry Matters (3rd Edition) Textbook — Section 7.3

Checkpoint 7.1

1. An unknown mass of silver nitrate, AgNO_3 , contains 7 g of nitrogen.
 - (a) Determine the percentage mass of nitrogen in silver nitrate.
 - (b) Hence, calculate the mass of silver nitrate present.
 - (c) Based on your answer in (b), determine the number of ions present in silver nitrate.
2. A student was provided with 1.5 mol of ammonia gas, NH_3 .
 - (a) Calculate the number of hydrogen atoms in 1.5 mol of ammonia gas.
 - (b) What is the mass of 1.5 mol of ammonia gas?
 - (c) What is the mass of one molecule of ammonia gas?
3. A sample of 0.5 mol of an unknown metal carbonate, MCO_3 , has a mass of 62.5 g.
 - (a) Calculate the molar mass of the metal carbonate.
 - (b) Based on your answer in (a), identify metal M.

D What Are Empirical and Molecular Formulae?

Learning Outcome

- Calculate empirical and molecular formulae from relevant data.

1. The **molecular formula** of a substance shows the exact number of atoms present in the substance.
2. For example, the molecular formula of carbon dioxide is CO_2 . This indicates that a carbon dioxide molecule is made up of one carbon atom and two oxygen atoms.
3. The **empirical formula** of a substance is the simplest form of the molecular formula. It shows the ratio of elements present in the formula in its simplest form.
4. Some examples of molecular formulae and their empirical formulae are shown below.

Molecular Formula	M_r of Molecular Formula	Empirical Formula	M_r of Empirical Formula
N_2H_4	32	NH_2	16
$\text{C}_2\text{H}_4\text{O}_2$	60	CH_2O	30
$\text{C}_6\text{H}_{12}\text{O}_6$	180	CH_2O	30
P_2O_5	220	P_2O_5	110

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