

## Chemistry Semester 2 Review

### Quarter Three Performance Expectations

#### Unit 10: The Mole

GOT (TII)	Task I should have mastered
	Relate the mole to common counting units such as dozens and pairs.
	Compare the relative sizes of Avogadro's number and common counting units.
	Identify the representative particles of elemental, ionic, and molecular substances.
	Calculate molar mass for compounds and molecules.
	Use molar mass and Avogadro's number to convert mass to moles and moles to mass for pure substances.
	Convert between moles to grams and moles to atoms for pure substances.

#### Unit 11: Chemical Reactions

GOT (TII)	Task I should have mastered
	Analyze chemical equations for evidence of a chemical reaction (production of a gas, precipitate, water, heat/light, or a new substance).
	Show evidence of the law of conservation of mass by writing a chemical equation in terms of moles, mass, or molecules.
	Represent solids (s), liquids (l), gases (g) and aqueous solutions (aq) in chemical reactions.
	Write word equations, skeleton equations, and balanced chemical equations.
	Classify common reactions as combustion, synthesis, decomposition, single replacement, or double replacement.
	Predict the products of hydrocarbon combustion reactions, write the equation, and balance it.
	Apply the mole concept and conservation of mass to chemical reactions.
	Distinguish and define limiting versus excess reactants in a chemical reaction.

#### Unit 12: Thermochemistry

GOT (TII)	Task I should have mastered
	Model energy changes in chemical reactions and relate them to the laws of conservation of energy and matter.
	Relate chemical potential energy to exothermic and endothermic reactions.
	Analyze calorimetry data to determine changes to the system, surroundings, and universe.
	Analyze a time/temperature graph for phase changes.
	Summarize the concepts of heat of vaporization, heat of fusion, and specific heat capacity.
	Predict the of enthalpy for exothermic and endothermic reactions.
	Cite evidence to explain how collision theory relates to reaction rates.
	Summarize the factors that affect reaction rates.

# Chemistry Semester 2 Course Review

**Stanford University**



## Chemistry Semester 2 Course Review:

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- Chapter 2: Essential Elements of Chemistry Semester 2 Course Review
- Chapter 3: Chemistry Semester 2 Course Review in Everyday Life
- Chapter 4: Chemistry Semester 2 Course Review in Specific Contexts
- Chapter 5: Conclusion

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6. In chapter 5, the author will draw a conclusion about Chemistry Semester 2 Course Review. The final chapter will summarize the key points that have been discussed throughout the book.

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## **Table of Contents Chemistry Semester 2 Course Review**

1. Understanding the eBook Chemistry Semester 2 Course Review
  - The Rise of Digital Reading Chemistry Semester 2 Course Review
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Semester 2 Course Review
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Semester 2 Course Review
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Semester 2 Course Review
  - Personalized Recommendations
  - Chemistry Semester 2 Course Review User Reviews and Ratings
  - Chemistry Semester 2 Course Review and Bestseller Lists
5. Accessing Chemistry Semester 2 Course Review Free and Paid eBooks
  - Chemistry Semester 2 Course Review Public Domain eBooks
  - Chemistry Semester 2 Course Review eBook Subscription Services
  - Chemistry Semester 2 Course Review Budget-Friendly Options
6. Navigating Chemistry Semester 2 Course Review eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Semester 2 Course Review Compatibility with Devices
  - Chemistry Semester 2 Course Review Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Semester 2 Course Review
  - Highlighting and Note-Taking Chemistry Semester 2 Course Review
  - Interactive Elements Chemistry Semester 2 Course Review
8. Staying Engaged with Chemistry Semester 2 Course Review

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Semester 2 Course Review
9. Balancing eBooks and Physical Books Chemistry Semester 2 Course Review
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Semester 2 Course Review
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Semester 2 Course Review
- Setting Reading Goals Chemistry Semester 2 Course Review
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Semester 2 Course Review
- Fact-Checking eBook Content of Chemistry Semester 2 Course Review
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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