

**Balancing a Redox Equation for an Acidic Solution
Using the Oxidation Number Method**
(Student textbook page 615)

35. Use the oxidation number method to balance the following equation for the combustion of carbon disulfide:



What Is Required?

You must balance an equation for the combustion of carbon disulfide, using the oxidation number method.

What Is Given?

You are given the reactants, carbon disulfide and oxygen, and the products, carbon dioxide and sulfur dioxide.

You are given rules for assigning oxidation numbers in **Table 9.3** on page 604 of the student textbook.

| Plan Your Strategy | Act on Your Strategy |
|--|---|
| Step 1 Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction. | $\overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g})$ <p>The oxidation numbers of sulfur and oxygen change during the reaction. It is a redox reaction.</p> |
| Step 2 Identify the atom or atoms that undergo an increase in oxidation number and the atom or atoms that undergo a decrease in oxidation number. | <p>Each of two sulfur atoms lost 6 electrons and each of two oxygen atoms gained 2 electrons. So the O_2 lost 12 electrons and the CS_2 gained 4 electrons.</p> $ \begin{array}{c} \text{1}(-12\text{e}^-) = -12\text{e}^- \\ \downarrow \\ \overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g}) \\ \uparrow \qquad \qquad \qquad \uparrow \qquad \qquad \uparrow \\ \text{3}(+4\text{e}^-) = +12\text{e}^- \end{array} $ |

Chemistry Solutions Manual Redox Reactions

Alen Hadzovic



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Chemistry in Context - Laboratory Manual Graham Hill, John Stranger Holman, John Holman, 2001 The laboratory manual and study guide supports your teaching with a broad range of practicals emphasising safety and risk assessment It is an essential companion to *Chemistry in Context* and can also be used alongside other Advanced Chemistry books It offers practicals with detailed instructions for openended investigations and opportunities for assessed practical work in the four skill areas of planning implementing analysing and evaluating **Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition** John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter

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additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

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Katherine J. Denniston,1995 This student study guide has been written to accompany Foundations of Inorganic Organic and Biological Chemistry by Caret Denniston and Topping It was designed to complement the text not to be used in place of the text Each chapter of the study guide contains the following sections reorganized and expanded set of learning goals concise chapter summary in chapter solved problems list of key terms review problems Preface

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