

Spring Final Review**MOLES**

1. The gram formula mass of Al_2O_3 is 101.961 g/mol. What is the mass of 6.02×10^{23} molecules of Al_2O_3 ?

$$101.961 \text{ g Al}_2\text{O}_3$$

2. How many molecules are in 9.45 mol of sodium nitrate, NaNO_3 ?

$$9.45 \text{ mol NaNO}_3 \times \frac{6.02 \times 10^{23} \text{ NaNO}_3}{1 \text{ mol NaNO}_3} = 5.69 \times 10^{24} \text{ NaNO}_3$$

3. How many grams are in 1.58 mol of potassium nitride, K_3N ?

$$1.58 \text{ mol K}_3\text{N} \times \frac{131.3 \text{ g K}_3\text{N}}{1 \text{ mol K}_3\text{N}} = 207 \text{ g K}_3\text{N}$$

4. How many molecules are in 3.47 g of MgCl_2 ?

$$3.47 \text{ g MgCl}_2 \times \frac{1 \text{ mol MgCl}_2}{95.21 \text{ g MgCl}_2} \times \frac{6.02 \times 10^{23} \text{ MgCl}_2}{1 \text{ mol MgCl}_2} = 2.19 \times 10^{22} \text{ MgCl}_2$$

5. Nicotine is 74.1% carbon, 8.6% hydrogen, and 17.3% nitrogen by mass. What is its empirical formula?

"Percent by mass, gets to mole, divide by smallest, multiply 'til whole"



6. Analysis of a compound gave 26.2% N, 7.5% H, 66.4% Cl. It's molecular mass is 160 g/mol. a. What is its empirical formula?



- b. What is its molecular formula?

$$\frac{160}{53.45} = 3 \quad \text{N}_3\text{H}_{12}\text{Cl}_3$$

7. What is the percentage composition of silver in silver nitrate (AgNO_3)?

$$\frac{107.87}{169.87} \times 100 = 63.5 \% \text{ Ag}$$

STOICHIOMETRY

1. Copper(II)Oxide decomposes to form solid copper and oxygen gas. If 95.4 g of copper(II) oxide decomposes, how many grams of oxygen are formed?



$$\frac{95.4 \text{ g CuO}}{1} \times \frac{1 \text{ mol CuO}}{79.6 \text{ g CuO}} \times \frac{1 \text{ mol O}_2}{2 \text{ mol CuO}} \times \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 19.2 \text{ g O}_2$$

2. Potassium hydroxide reacts with sulfuric acid in a double replacement reaction. What is the mass of water if 19.6 g of KOH reacts with excess acid?



$$\frac{19.6 \text{ g KOH}}{1} \times \frac{1 \text{ mol KOH}}{56.1 \text{ g KOH}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol KOH}} \times \frac{18.0 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 3.14 \text{ g H}_2\text{O}$$

3. If 36.5 g of HCl and 73.0 g of Zn are put together, and the reaction proceeds as follows:



- a. Find the limiting reactant.

$$\frac{36.5 \text{ g HCl}}{1} \times \frac{1 \text{ mol HCl}}{36.5 \text{ g HCl}} \times \frac{1 \text{ mol ZnCl}_2}{2 \text{ mol HCl}} \times \frac{136.4 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

$$\frac{73.0 \text{ g Zn}}{1} \times \frac{1 \text{ mol Zn}}{65.4 \text{ g Zn}} \times \frac{1 \text{ mol ZnCl}_2}{1 \text{ mol Zn}} \times \frac{136.4 \text{ g ZnCl}_2}{1 \text{ mol ZnCl}_2}$$

36.5 g **HCl** is the limiting reactant (produces less ZnCl_2)

- b. Find the mass of ZnCl_2 formed.

36.5 g HCl produces 68.2 g ZnCl_2

73.0 g Zn produces 152 g ZnCl_2

- c. If only 80 g of ZnCl_2 is formed find the % yield. (Actual yield / theoretical yield) $\times 100$

$$(80 \text{ g} / 68.2 \text{ g}) \times 100 = 117\%$$

Chemistry Spring Final Review Answers

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