

Chemistry 338 2nd Semester Final Exam Review Sheet

District Assessment

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten	¹⁸⁴ W ⁺² 74	74	74	110	72	184	+2

- Name the compound
 a. Fe₂O₃ – **iron (III) oxide**
 b. Cl₂O₈ – **dichlorine octaoxide**
- Write the formula
 a. Trisulfur Dinitride - **S₃N₂**
 b. Zinc Hydroxide – **Zn(OH)₂**
 c. Tin (IV) oxide - **SnO₂**
- The pressure of a gas is 843 mm Hg at 24°C. What will the pressure be at standard temperature? Volume is constant.

$$\frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{843 \text{ mmHg}}{297 \text{ K}} = \frac{P_2}{273 \text{ K}} \quad P_2 = 775 \text{ mm Hg}$$
- The volume of a gas is 81.7 mL when the pressure is 119 kPa and the temperature is 73°C. What is the volume at STP?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \frac{(119 \text{ kPa})(81.7 \text{ mL})}{346 \text{ K}} = \frac{(101.3 \text{ kPa})(V_2)}{273 \text{ K}} \quad V_2 = 75.7 \text{ mL}$$
- Between Be and Ba, which has a larger ionization energy? Larger atomic radius?
Larger ionization energy – Be
Larger atomic radius - Ba
- Balance the equation & classify the type of reaction:

$$2\text{Al} + 3 \text{CuSO}_4 \rightarrow 1 \text{Al}_2(\text{SO}_4)_3 + 3 \text{Cu} \quad \text{single replacement}$$
- Predict the products & classify the type of reaction:

$$\text{H}_3\text{PO}_4 + 3 \text{NaOH} \rightarrow 3 \text{H}_2\text{O} + \text{Na}_3\text{PO}_4 \quad \text{double replacement}$$
- Calculate the Molar Mass: (NH₄)₂CO₃
96 g/mole
- Convert 2.75 moles of N₂O₅ to molecules

$$\frac{2.75 \text{ moles}}{1} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} = 1.66 \times 10^{24} \text{ molecules N}_2\text{O}_5$$

Solutions

Chemistry Spring Semester Final Exam Review Sheet

Victor M. Corman



Chemistry Spring Semester Final Exam Review Sheet:

Survival Handbook for the New Chemistry Instructor Diane M. Bunce, Cinzia M. Muzzi, 2004 This book provides an overview of the issues facing new chemistry faculty in preparation for teaching Serving as a reference to answer specific questions new chemistry faculty encounter this book is comparable to sitting down with a colleague in the department and talking through some ideas or gaining some pointers on how to avoid common pitfalls It is the one single place new chemistry faculty can go to find practical information on how to teach and how to prepare for teaching their first course Chapters are written both by established experts in the field and by new professors within their first couple of years of teaching

Cornell University Courses of Study Cornell University, 1992 **Congressional Record** United States.

Congress, 1971 The Congressional Record is the official record of the proceedings and debates of the United States Congress It is published daily when Congress is in session The Congressional Record began publication in 1873 Debates for sessions prior to 1873 are recorded in The Debates and Proceedings in the Congress of the United States 1789 1824 the Register of Debates in Congress 1824 1837 and the Congressional Globe 1833 1873 **The Journal of Materials Education**, 1998

Getting Into Medical School Donald J. Solomon, 1979 Virginia Journal of Science Ruskin Skidmore Freer, 2006

The Virginia Journal of Science, 2005

Popular Science, 2005-09 Popular Science gives our readers the information and tools to improve their technology and their world The core belief that Popular Science and our readers share The future is going to be better and science and technology are the driving forces that will help make it better **Popular Science**, 2004-12 Popular Science gives our readers the information and tools to improve their technology and their world The core belief that Popular Science and our readers share The future is going to be better and science and technology are the driving forces that will help make it better *Final Exam Review Sheet*, 2002 The Rural Educator, 2013 **Fast Track: Chemistry**

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