

Name: KEY Date: _____ Hr: _____

Stoichiometry worksheet #2

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1. The equation for the synthesis of ammonia is below. How many moles of H_2 are needed to produce 9.33 moles of NH_3 ?



$$\frac{9.33 \text{ mol } NH_3}{1} \times \frac{3 \text{ mol } H_2}{2 \text{ mol } NH_3} = \boxed{13.995 \text{ mol } H_2}$$

2. What mass (in grams) of copper is produced when 3.80 g of iron reacts with excess $CuSO_4$?



$$\frac{3.80 \text{ g } Fe}{1} \times \frac{1 \text{ mol } Fe}{55.845 \text{ g } Fe} \times \frac{1 \text{ mol } Cu}{1 \text{ mol } Fe} \times \frac{63.546 \text{ g } Cu}{63.546 \text{ g } Cu} = \boxed{4.32 \text{ g } Cu}$$

3. If 5.00g of CH_4 are burned according to the equation below, what mass in grams of CO_2 are produced?



$$\frac{5.00 \text{ g } CH_4}{1} \times \frac{1 \text{ mol } CH_4}{16.043 \text{ g } CH_4} \times \frac{1 \text{ mol } CO_2}{1 \text{ mol } CH_4} \times \frac{44.01 \text{ g } CO_2}{44.01 \text{ g } CO_2} = \boxed{13.7 \text{ g } CO_2}$$

4. In the reaction that follows, determine how many grams of aluminum chloride ($AlCl_3$) will be produced from 92.01 g of Cl_2 .



$$\frac{92.01 \text{ g } Cl_2}{1} \times \frac{1 \text{ mol } Cl_2}{70.906 \text{ g } Cl_2} \times \frac{2 \text{ mol } AlCl_3}{3 \text{ mol } Cl_2} \times \frac{133.34 \text{ g } AlCl_3}{1 \text{ mol } AlCl_3} = \boxed{115.4 \text{ g } AlCl_3}$$

Chemistry Stoichiometry Problem Sheet 2 Answers

Ronald A. Hites, Jonathan D. Raff



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 Relevant Problems for Chemical Principles Ian S. Butler,Arthur E. Grosser,1984

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