1. The combustion of a sample of butane, C_4H_{10} (lighter fluid), produced 2.46 grams of water.

- (a) How many moles of water formed?
- (b) How many moles of butane burned?
- (c) How many grams of butane burned?
- (d) How much oxygen was used up in moles?
- (e) How much oxygen was used up in grams?
- 2. Sodium metal burns in air according to the balanced reaction below.

- (a) How many moles of oxygen are needed to react with 9.5 g of sodium completely?
- (b) How many grams of sodium are needed to produce 12.5 g of sodium oxide?
- 3. In the combustion of 54.5 g of butane (C_4H_6), how many grams of CO_2 are produced? Write and balance the equation before solving.

Chemistry Stoichiometry Problem Sheet 3 Answers

James E. Brady, Fred Senese

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