

Chemistry ch 14-solutions

1. Solution-
2. Water and salt water— water travels from its beaker to salt water because less concentration to higher concentration
3. Solvent- high in volume and can dissolve solute particles
4. Solute- less volume can be dissolved by solvent
 1. Liquid-liquid and gas-liquid forms are important
5. If solvent and solute form a solution, they lose potential energy
 1. Solvent-solvent (E1) — — break interactions... endothermic
 2. Solute- solute (E2) — — — break interactions need to overcome both of them to form a solution... endothermic
 3. Solvent-solute (E3) — — solution. STRONGEST interactions — less energy more stable.... exothermic
6. Solvent-solute > 1+2 = solution forms
7. Solvent-solute = 1+2 = solution forms
8. Solvent-solute < 1+2 = solution may or may not form based on relative disparity
9. Homogenous mixture- the look similar— UNIFORM
10. Entropy- energy spread — — even if 8 occurs, solution can occur because of entropy.
 1. Disorder- mixing
 2. C-C+ X-wrong unstable
 3. S+C — CS+ ==V stable
11. Atoms want to loose energy or decrease it as much as they can . Charged atoms have higher energy unstable
12. Increased entropy = decreasing energy
13. Ideal gases mix because mixing increases their entropy — — when it comes to gases it is always entropy not abt potential energy
14. Like dissolves like
 1. Polar molecules and ionic are more stable in polar solutions.
 2. Non polar molecules dissolve in non polar substances
15. Carbons don't like to dissolve in water
16. Heat of solution —
 1. Container releases energy- — gets hot
 2. Breaking solute $\Delta H > 0$ because it is endothermic— takes energy
 3. Breaking solvent $\Delta H > 0$ because endothermic so needs energy to break bonds
 4. Form solvent + solute = exothermic because releases energy
 5. If ΔH solute + ΔH solvent < ΔH solution = exothermic
 6. If ΔH solute + ΔH solvent > ΔH solution = endothermic
17. Heats Of Hydration
 1. Enthalpy change (ΔH) when 1 mol of gaseous ions is dissolved in water.
 2. $KI(s) \longrightarrow K^+(g) + I^-(g) \longrightarrow$ dissolves in water $\longrightarrow K^+(aq) + I^-(aq)$ solution

Chemistry Ch 14 Test Answer Key

George Rudolph Herrmann



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