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CHAPTER 8 REVIEW

## Chemical Equations and Reactions

#### SHORT ANSWER Answer the following questions in the space provided.

According to the law of conservation of mass, the total mass of the reacting substances is

- (a) always more than the total mass of the products.
- (b) always less than the total mass of the products.
- (c) sometimes more and sometimes less than the total mass of the products.
- (d) always equal to the total mass of the products.

## In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

Dute:

each question.		
<ol> <li>In the reaction described by the word</li> </ol>	equation	
sodium + ·	water sodium h	ydroxide + hydrogen gas
<ul> <li>a. sodium hydroxide is a product.</li> </ul>	b.	sodium is a product.
c. water is a reactant.	d.	Both (a) and (c)
3. Which of the following indicates that	a chemical equation	on is balanced?
a. The numbers of atoms of each ele-	ment are the same	on both sides of the equation.
<ul> <li>All of the coefficients are the same</li> </ul>	e.	
c. The numbers of molecules on each	h side of the equati	ion are equal.
<li>d. The sums of the coefficients on ea</li>	sch side of the equi	ation are equal.
<ol> <li>From a complete and correctly writter</li> </ol>	n chemical equatio	en, you can obtain the
<ul> <li>a. chemical formulas of the reactants</li> </ul>		
<ul> <li>b. relative amounts of the reactants a</li> </ul>		
<ul> <li>physical states of the reactants and</li> </ul>	d products.	
d. All of the above		
5. To balance a chemical equation, you	must	
<ul> <li>a. adjust the subscripts.</li> </ul>	lb.	adjust the number of products.
<ul> <li>c. adjust the number of reactants.</li> </ul>	d.	adjust the coefficients.
6. The formation of a solid product in a		
a. (aq). b. (g).	e. (/).	d. (s).
8. In the expression 3CO <sub>2</sub> , the numbers .	3 and 2 are, respec-	
<ul> <li>a. a subscript and a coefficient.</li> </ul>	D	a coefficient and a subscript.
c. two subscripts.	d.	two coefficients.
9. What is the total number of molecules	s that is represente NHaNO <sub>2</sub> > N	d by the following chemical equation? (Add up the coefficients.)
a. 3 b. 4	c. 8	d. 16
10. Which equation below violates the lav		
a. 2H₂ + O₂ → 2H₂O	D.	$KCI + Br \rightarrow KBr + Cb$
c. 2Fe <sub>2</sub> O <sub>2</sub> + 3C → 4Fe + 3CO <sub>2</sub>	d.	$Na_2CO_1 + 2HC1 \rightarrow 2NaC1 + H_2O + CO_2$
		ed equation for the reaction of iron and oxygen?
a. 2Fe + O <sub>2</sub> → Fe <sub>2</sub> O <sub>1</sub>	ь.	$Fe + 3O_2 \rightarrow Fe_2O_3$
e. 4Fe + 3O <sub>2</sub> → 2Fe <sub>2</sub> O <sub>1</sub>	-4	$3Fe + 3O_2 \rightarrow 2Fe_2O_2$
12. Which coefficients correctly balance t	the formula equati-	
	$KClO_2(x) \rightarrow KC$	
a, 1, 1, 1 b, 1, 1, 3	c. 2, 2, 3	d. 2, 1, 1
13. Which equation is not balanced?		
a. $2H_2 + O_2 \rightarrow 2H_2O$	b.	$4H_2 + 2O_2 \rightarrow 4H_2O$
c. H <sub>2</sub> + H <sub>2</sub> + O <sub>2</sub> → H <sub>2</sub> O + H <sub>2</sub> O	at.	$2H_2 + O_2 \rightarrow H_2O$
14. In which type of reaction do two or m	sore compounds re	
a. synthesis (AKA combination)	ъ.	decomposition
c. combustion	d.	single-replacement
15. Which of the following is true regards	ing single-replacer	
<ul> <li>a. Only one product is formed.</li> </ul>	ъ.	There is only one reagent.
<ul> <li>a. It requires oxygen.</li> </ul>	d.	One element replaces another in a compound.

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Peter P. Berlow

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