AP Chemistry— CHAPTER 16 STUDY GUIDE— Acid-Base Equilibrium

16.1 Acids and Bases: A Brief Review

- Acids taste sour and cause certain dyes to change color.
- ·Bases taste bitter and feel soapy.
- Arrhenius concept of acids and bases:
 - ·An acid is a substance that, when dissolved in water, increases the concentration of H ions.
 - Example: HCl is an acid.
 - An Arrhenius base is a substance that, when dissolved in water, increases the concentration of OH ions.
 - ·Example: NaOH is a base.
 - This definition is quite narrow in scope as it limits us to aqueous solutions.

16.2 Brønsted-Lowry Acids and Bases

 We can use a broader, more general definition for acids and bases that is based on the fact that acid-base reactions involve proton transfers.

The H Ion in Water

- The H^{*}(ag) ion is simply a proton with no surrounding valence electrons.
- ·In water, clusters of hydrated H (aa) ions form.
- The simplest cluster is H₁O (aq).
 - We call this a hydronium ion.
 - Larger clusters are also possible (such as H₂O₂" and H₂O₄").
- Generally we use H'(aq) and H₂O'(aq) interchangeably.

Proton-Transfer Reactions

- ·We will focus our attention on H (aq).
- According to the Arrhenius definitions, an acid increases [H] and a base increases [OH].
- Another definition of acids and bases was proposed by Bronsted and Lowry.
- In the Bronsted-Lowry system, a Bronsted-Lowry acid is a species that donates H* and a Bronsted-Lowry base is a species that accepts H*.
 - Therefore a Bronsted-Lowry base does not need to contain OH.
 - *NH₃ is a Bronsted-Lowry base, but not an Arrhenius base.
- •Consider $NH_1(ag) + H_2O(I) \rightarrow NH_1(ag) + OH(ag)$:
 - ·H₂O donates a proton to ammonia.
 - Therefore, water is acting as an acid.
 - NH₅ accepts a proton from water.
 - ·Therefore, ammonia is acting as a base.
 - An amphiprotic substance can behave either as an acid or as a base.
 - Thus, water is an example of an amphiprotic species.

Conjugate Acid-Base Pairs

- Whatever is left of the acid after the proton is donated is called its conjugate base.
- Similarly, a conjugate acid is formed by adding a proton to the base.
- •Consider $HX(aq) + H_2O(l) \rightarrow H_2O'(aq) + X'(aq)$:
 - ·HX and X differ only in the presence or absence of a proton.
 - They are said to be a conjugate acid-base pair.
 - X is called the conjugate base.
 - After HX (acid) loses its proton it is converted into X⁻ (base).
 - *Therefore HX and X are a conjugate acid-base pair.
 - After H₂O (base) gains a proton it is converted into H₂O' (acid).
 - ·H.O is the conjugate acid.
 - *Therefore, H₂O and H₃O' are a conjugate acid-base pair.

Relative Strengths of Acids and Bases

- ·The stronger an acid is, the weaker its conjugate base will be.
- We can categorize acids and bases according to their behavior in water.
 - 1. Strong acids completely transfer their protons to water.
 - No undissociated molecules remain in solution.
 - Their conjugate bases have negligible tendencies to become protonated.
 - An example is HCl.

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