

# Chapter 18

## Study Guide

- Write the expression for the equilibrium constant for each of the following reactions.
  - $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2\text{HCl}(\text{g})$
  - $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightleftharpoons 2 \text{NH}_3(\text{g})$
  - $\text{HCN}(\text{g}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{CN}^-(\text{aq})$
  - $\text{H}_2\text{SO}_4(\text{g}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{HSO}_4^-(\text{aq})$
  - $\text{PbF}_2(\text{s}) \rightleftharpoons \text{Pb}^{2+}(\text{aq}) + 2\text{F}^-(\text{aq})$
- From the data provided below, calculate the value of the equilibrium constant for the reaction.
  - $\text{H}_2(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons 2\text{HCl}(\text{g})$   
 $[\text{H}_2] = [\text{Cl}_2] = 1.0 \times 10^{-2}$ ;  $[\text{HCl}] = 1.0 \times 10^{-4}$
  - $\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightleftharpoons 2 \text{NH}_3(\text{g})$   
 $[\text{N}_2] = 4.4 \times 10^{-2}$ ;  $[\text{H}_2] = 1.2 \times 10^{-1}$ ;  $[\text{NH}_3] = 3.4 \times 10^{-3}$
- For the reaction  $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$ ,  $K_{\text{eq}} = 0.2$ . At a particular time, the following concentrations are measured:  $[\text{N}_2\text{O}_4] = 2.0 \text{ M}$ ,  $[\text{NO}_2] = 0.2 \text{ M}$ . Is this reaction at equilibrium? If not, in which direction will the reaction proceed?
- For the reaction  $2\text{ICl}(\text{g}) \rightleftharpoons \text{I}_2(\text{g}) + \text{Cl}_2(\text{g})$ ,  $K_{\text{eq}} = 0.11$ . At a particular time, the following concentration are measured:  $[\text{ICl}] = 2.5 \text{ M}$ ,  $[\text{I}_2] = 2.0 \text{ M}$ ,  $[\text{Cl}_2] = 1.2 \text{ M}$ . Is this reaction at equilibrium? If not, in which direction will the reaction proceed?
- Name the 3 stresses that may be applied to a chemical equilibrium to cause it to shift.
- Describe LeChâtelier's Principle.
- For the reaction below, mark whether the stress listed will cause the reaction to move forward or in reverse.



- |                            |                               |                            |
|----------------------------|-------------------------------|----------------------------|
| a. Increase Heat           | e. Increase $[\text{ZnCl}_2]$ | i. Decrease $[\text{H}_2]$ |
| b. Increase Pressure       | f. Increase $[\text{HCl}]$    | j. Decrease Pressure       |
| c. Increase $[\text{H}_2]$ | g. Decrease $[\text{HCl}]$    | k. Decrease Heat           |
| d. Increase $[\text{Zn}]$  | h. Decrease $[\text{ZnCl}_2]$ |                            |

- In a system at equilibrium, the \_\_\_\_\_ are equal.
- In a system at equilibrium, the \_\_\_\_\_ are constant.
- $K_{\text{sp}}$  depends on the \_\_\_\_\_ of the solution.
- How many ions are produced by the dissolution of a unit of  $\text{CaCl}_2$ ?
- When referred to in terms of  $K_{\text{sp}}$ , Q is called the \_\_\_\_\_. If  $Q > K_{\text{sp}}$ , the solution is \_\_\_\_\_. If  $Q < K_{\text{sp}}$ , the solution is \_\_\_\_\_. If  $Q = K_{\text{sp}}$ , the solution is \_\_\_\_\_.
- What is the common-ion effect?
- Write the expression for the solubility product constant for  $\text{SrSO}_4$ .
- Write the expression for the solubility product constant for  $\text{Al}_2(\text{SO}_4)_3$ .
- Write the expression for the solubility product constant for  $\text{AgI}$ .
- A sample of  $\text{SrCO}_3(\text{s})$  is added to pure water and allowed to come to equilibrium at  $25^\circ\text{C}$ . The concentration of  $\text{Sr}^{2+}$  is  $4.0 \times 10^{-5} \text{ M}$  at equilibrium. What is the value of  $K_{\text{sp}}$  for  $\text{SrCO}_3$ ?

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**John Krenos, Joseph Potenza, Loretta  
Jones, Lynn Koplitz, Thomas Spence**



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### **Table of Contents Chemistry Chapter 18 Study Guide**

1. Understanding the eBook Chemistry Chapter 18 Study Guide
  - The Rise of Digital Reading Chemistry Chapter 18 Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 18 Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 18 Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 18 Study Guide
  - Personalized Recommendations
  - Chemistry Chapter 18 Study Guide User Reviews and Ratings
  - Chemistry Chapter 18 Study Guide and Bestseller Lists

5. Accessing Chemistry Chapter 18 Study Guide Free and Paid eBooks
  - Chemistry Chapter 18 Study Guide Public Domain eBooks
  - Chemistry Chapter 18 Study Guide eBook Subscription Services
  - Chemistry Chapter 18 Study Guide Budget-Friendly Options
6. Navigating Chemistry Chapter 18 Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Chapter 18 Study Guide Compatibility with Devices
  - Chemistry Chapter 18 Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 18 Study Guide
  - Highlighting and Note-Taking Chemistry Chapter 18 Study Guide
  - Interactive Elements Chemistry Chapter 18 Study Guide
8. Staying Engaged with Chemistry Chapter 18 Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 18 Study Guide
9. Balancing eBooks and Physical Books Chemistry Chapter 18 Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 18 Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Chapter 18 Study Guide
  - Setting Reading Goals Chemistry Chapter 18 Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Chapter 18 Study Guide
  - Fact-Checking eBook Content of Chemistry Chapter 18 Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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