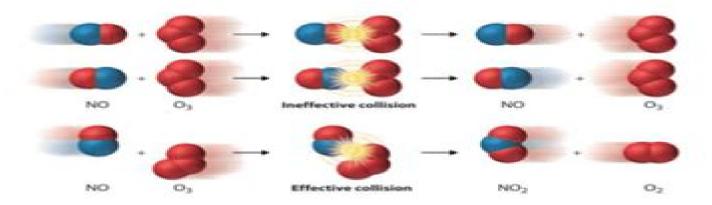
Topic	Collision theory		Level	GCSE
Outcomes	70.00	reaction can concentratio	increase the n, surface ar	y to explain why changes to a rate of reaction (temperature, ea and catalyst) Il collisions are successful

The collision theory and rates of reaction



When two particles collide (hit each other) sometimes a chemical reaction happens. Not all collisions result in a chemical reaction. To increase the number of successful collisions we can change the conditions of a reaction (see page 2).

- 1. Do all collisions between particles cause a reaction to happen?
- What do we call the minimum energy the particles must have for a reaction to happen?
- 3. What do we call a collision between two particles that causes a chemical reaction?
- Look at the picture above, how do you know the first reaction is not a successful collision?

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