

## 12U Equilibrium Unit Test

Full Name: \_\_\_\_\_

Date: \_\_\_\_\_

### Multiple Choice K/U

Identify the letter of the choice that best completes the statement or answers the question.

- \_\_\_\_\_ 1. When solid lead(II) phosphate is in equilibrium with its ions, the ratio of lead(II) ions to phosphate ions is which of the following?
- 1:1
  - 1:2
  - 2:1
  - 2:3
  - 3:2
- \_\_\_\_\_ 2. In a saturated solution of silver phosphate, the concentration of silver ion is  $4.5 \times 10^{-4}$  mol/L. The  $K_{sp}$  of silver phosphate would be which of the following?
- $6.8 \times 10^{-8}$
  - $1.4 \times 10^{-14}$
  - $1.5 \times 10^{-13}$
  - $1.0 \times 10^{-13}$
  - none of the above
- \_\_\_\_\_ 3. For the equilibrium system below, which of the following would result in an increase in the quantity of  $\text{PCl}_3(\text{g})$ ?
- $$\text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{PCl}_5(\text{g}) + 45 \text{ kJ}$$
- increasing temperature
  - increasing the size of the container
  - decreasing temperature
  - removing some  $\text{Cl}_2(\text{g})$
  - injecting some He gas
- \_\_\_\_\_ 4. 10.0 mol of ammonia gas is injected into a 4.0 L container. At equilibrium 1.2 mol of hydrogen gas is found in the container. The number of moles of ammonia gas left in the container must be which of the following?
- $$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g}) + 94 \text{ kJ}$$
- 6.8
  - 6.4
  - 8.6
  - 8.8
  - 9.2
- \_\_\_\_\_ 5. A closed flask contains this equilibrium
- $$2\text{NO}_2(\text{g}) \text{ (brown gas)} \rightleftharpoons \text{N}_2\text{O}_4(\text{g}) \text{ (colourless gas)}$$
- When the equilibrium is heated at a constant pressure it gets browner. This must mean which of the following?
- $K_c$  is very large
  - it is endothermic
  - it is exothermic
  - both a and b
  - both a and c
- \_\_\_\_\_ 6. Which equilibrium shows a decrease in [products] when the volume increases?
- $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g})$
  - $\text{Cl}_2(\text{g}) + \text{PCl}_3(\text{g}) \rightleftharpoons \text{PCl}_5(\text{g})$
  - $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
  - $2\text{NH}_3(\text{g}) \rightleftharpoons \text{N}_2 + 3\text{H}_2(\text{g})$
- I
  - I and II
  - IV
  - III and IV
  - II
- \_\_\_\_\_ 7. Consider this equilibrium
- $$4\text{HCl}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g}) + 2\text{Cl}_2(\text{g})$$
- The equilibrium law expression for the balanced chemical equation would be
- $[\text{HCl}][\text{O}_2] / [\text{H}_2\text{O}][\text{Cl}_2]$
  - $[\text{H}_2\text{O}]^2[\text{Cl}_2]^2 / [\text{HCl}]^4[\text{O}_2]$
  - $2[\text{H}_2\text{O}][\text{Cl}_2] / 4[\text{HCl}][\text{O}_2]$
  - $[\text{HCl}]^4[\text{O}_2] / [\text{H}_2\text{O}]^2[\text{Cl}_2]^2$
  - $[\text{H}_2\text{O}][\text{Cl}_2] / \text{HCl}[\text{O}_2]$
- \_\_\_\_\_ 8. The  $K_{sp}$  for a substance  $\text{A}_2\text{B}$  is  $1.2 \times 10^{-12}$ . The  $[\text{B}]$  must be which of the following?
- $4.0 \times 10^{-11}$
  - $3.6 \times 10^{-12}$
  - $6.7 \times 10^{-5}$
  - $1.3 \times 10^{-4}$
  - $2.6 \times 10^{-4}$

# Chemistry Equilibrium Multiple Choice Practice Test

**Mary Millhollon, Richard H. Langley**



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