

- B) H_2S is a larger molecule, leading to stronger London dispersion forces than for H_2O
- C) water has a low mass
- D) hydrogen bonds exist among molecules of H_2O
- E) H_2O is a larger molecule, leading to stronger London dispersion forces than for H_2S

80. Convert the pressure 1.30 atm to kPa.

- A) 2 kPa B) 115 kPa C) 132 kPa D) 245 kPa

81. Which of these changes would NOT cause an increase in the pressure of a contained gas?

- A) The volume of the container is increased
- B) More of the gas is added to the container
- C) The temperature is increased
- D) The average kinetic energy of the gas is increased

82. What is the volume, in liters, occupied by 1.73 moles of N_2 gas at 0.992 atm pressure and a temperature of 75°C ?

- A) 10.7 L B) 33.8 L C) 49.8 L D) 52.2 L

83. Under ideal conditions, the molar volume of He gas at STP is 22.4 L/mol and the molar volume of Ne gas at SATP is 24.8 L/mol. Which one of the following statements is correct?

- A) The number of molecules in both volumes is 6.02×10^{23} .
- B) There are more molecules in the 24.8 L of Ne than in the 22.4 L of He.
- C) The mass of the He sample is greater than the mass of the Ne sample.
- D) The number of molecules cannot be compared for two different gases.

84. The density of N_2O at 1.53 atm and 45.2°C is _____ g/L.

- A) 18.2 B) 1.76 C) 0.388 D) 9.99 E) 2.58

85. The reaction of 50 mL of Cl_2 gas with 50 mL of CH_4 gas via the equation:
 $\text{Cl}_{2(g)} + \text{CH}_{4(g)} \rightarrow \text{HCl}_{(g)} + \text{CH}_3\text{Cl}_{(g)}$
 will produce a total of _____ mL of products if pressure and temperature are kept constant.

- A) 100 B) 50 C) 200 D) 150 E) 250

86. If a gas has an initial volume of 10.0 L at 1.0 atm and 350 K, what will the volume be at 2.0 atm and 700 K?

- A) 5.0 L
- B) 10.0 L
- C) 20.0 L
- D) 30.0 L

87. Which is NOT true according to the kinetic molecular theory of gases?

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