



Rules for Writing the Expression for Equilibrium Constant for Heterogeneous Reaction

- The concentration terms of products are written in the numerator and the concentration terms of reactants are written in the denominator raised to the power of their respective coefficient in the balanced chemical equation.
- While writing the expression for K_c , the molarities of gaseous reactants and products and species present in the solution are written.
- Concentration terms of solids and pure liquids do not appear in the expression, as they are taken to be unity.
- While writing the expression for K_p , partial pressures of only the gaseous reactants and products are written in the expression for equilibrium constant.

Example 5.3

The equilibrium constant expression for some heterogeneous reactions are shown below:



$$K_p = p_{\text{CO}_2}$$



$$K_p = (p_{\text{NH}_3})(p_{\text{HCl}})$$



$$K_c = \frac{[\text{H}^+][\text{HCO}_3^-]}{[\text{CO}_2]}$$



$$K_c = \frac{1}{[\text{Fe}^{3+}][\text{OH}^-]^3}$$

Exercise 5.4

Answer the following questions:

1. 5.0 mole of ammonia were introduced into a 5.0 L reaction chamber in which it is partially decomposed at high temperatures.

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