

Q.1. The bond length and bond angle in the molecules of CH_4 , NH_3 and H_2O are given below:



The variation of bond angles is a result of

1. increasing repulsion between H atoms as bond length decreases.
2. Number of non bonding electron pairs in the molecule
3. non-bonding electron pair has greater repulsive force than bonding electron pair.

(A) 1, 2, 3 (B) only 1 & 2 (C) only 2 & 3 (D) only 1 -

Q.2. The lattice enthalpy and hydration enthalpy of four compounds are given below

Compound	Lattice enthalpy (KJ/mol)	Hydration enthalpy (KJ/mol)
P	+780	-920
Q	+1012	-812
R	+828	-578
S	+632	-600

The pair of compounds soluble in water is

(A) P, R (B) Q, R (C) R, S (D) Q, S

3. The root mean square speed of the molecules of a diatomic gas is 'u'. When temperature is doubled, the molecules dissociate into two atoms. The new rms speed of the atom is

(A) $\sqrt{2}u$ (B) u (C) 2u (D) 4u

4. The melting point of RbBr is 682°C while that of NaF is 988°C . The much higher m.p. of NaF is mainly due to

- (A) the two crystals are not isomorphous
- (B) molar mass of NaF is smaller than that of RbBr
- (C) internuclear distance of RbBr is greater than that of NaF
- (D) bond in RbBr is more covalent than the bond in NaF

5. Solubility product constants (K_{sp}) of salts of type MX , MX_2 and M_3X at temperature T are 4×10^{-8} , 3.2×10^{-14} and 2.7×10^{-15} respectively. Solubility (mol dm^{-3}) of the salts at temperature T are in the order:

- (A) $\text{MX} > \text{MX}_2 > \text{M}_3\text{X}$
- (B) $\text{M}_3\text{X} > \text{MX}_2 > \text{MX}$
- (C) $\text{MX}_2 > \text{M}_3\text{X} > \text{MX}$
- (D) $\text{MX} > \text{M}_3\text{X} > \text{MX}_2$

6. f orbital has

- (A) zero radial node
- (B) two total nodes
- (C) two angular nodes
- (D) all of these

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