

## Chemistry Semester 2 Review

### Quarter Three Performance Expectations

#### Unit 10: The Mole

GOT (TII)	Task I should have mastered
	Relate the mole to common counting units such as dozens and pairs.
	Compare the relative sizes of Avogadro's number and common counting units.
	Identify the representative particles of elemental, ionic, and molecular substances.
	Calculate molar mass for compounds and molecules.
	Use molar mass and Avogadro's number to convert mass to moles and moles to mass for pure substances.
	Convert between moles to grams and moles to atoms for pure substances.

#### Unit 11: Chemical Reactions

GOT (TII)	Task I should have mastered
	Analyze chemical equations for evidence of a chemical reaction (production of a gas, precipitate, water, heat/light, or a new substance).
	Show evidence of the law of conservation of mass by writing a chemical equation in terms of moles, mass, or molecules.
	Represent solids (s), liquids (l), gases (g) and aqueous solutions (aq) in chemical reactions.
	Write word equations, skeleton equations, and balanced chemical equations.
	Classify common reactions as combustion, synthesis, decomposition, single replacement, or double replacement.
	Predict the products of hydrocarbon combustion reactions, write the equation, and balance it.
	Apply the mole concept and conservation of mass to chemical reactions.
	Distinguish and define limiting versus excess reactants in a chemical reaction.

#### Unit 12: Thermochemistry

GOT (TII)	Task I should have mastered
	Model energy changes in chemical reactions and relate them to the laws of conservation of energy and matter.
	Relate chemical potential energy to exothermic and endothermic reactions.
	Analyze calorimetry data to determine changes to the system, surroundings, and universe.
	Analyze a time/temperature graph for phase changes.
	Summarize the concepts of heat of vaporization, heat of fusion, and specific heat capacity.
	Predict the of enthalpy for exothermic and endothermic reactions.
	Cite evidence to explain how collision theory relates to reaction rates.
	Summarize the factors that affect reaction rates.

# Chemistry Semester Two Course Review

**R Pring**



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### **Table of Contents Chemistry Semester Two Course Review**

1. Understanding the eBook Chemistry Semester Two Course Review
  - The Rise of Digital Reading Chemistry Semester Two Course Review
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Semester Two Course Review
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Semester Two Course Review
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Semester Two Course Review
  - Personalized Recommendations
  - Chemistry Semester Two Course Review User Reviews and Ratings
  - Chemistry Semester Two Course Review and Bestseller Lists
5. Accessing Chemistry Semester Two Course Review Free and Paid eBooks

- Chemistry Semester Two Course Review Public Domain eBooks
- Chemistry Semester Two Course Review eBook Subscription Services
- Chemistry Semester Two Course Review Budget-Friendly Options
- 6. Navigating Chemistry Semester Two Course Review eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Semester Two Course Review Compatibility with Devices
  - Chemistry Semester Two Course Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Semester Two Course Review
  - Highlighting and Note-Taking Chemistry Semester Two Course Review
  - Interactive Elements Chemistry Semester Two Course Review
- 8. Staying Engaged with Chemistry Semester Two Course Review
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Semester Two Course Review
- 9. Balancing eBooks and Physical Books Chemistry Semester Two Course Review
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Semester Two Course Review
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Semester Two Course Review
  - Setting Reading Goals Chemistry Semester Two Course Review
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Semester Two Course Review
  - Fact-Checking eBook Content of Chemistry Semester Two Course Review
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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