

Solve each of the following problems. Use correct units, and show your work for full credit.

1. The specific heat of ethanol is $2.46 \text{ J/g}^\circ\text{C}$. Find the heat energy required to raise the temperature of 193 g of ethanol from 19°C to 35°C .

$$q = (193)(2.46)(35-19)$$

$$\boxed{q = 7596 \text{ J}}$$

2. When a 120 g sample of aluminum absorbs 9612 J of heat energy, its temperature increases from 25°C to 115°C . Find the specific heat of aluminum. Include the correct unit.

$$C = \frac{q}{m \Delta T} = \frac{9612}{(120)(115-25)} \quad \boxed{0.89 \text{ J/g}^\circ\text{C}}$$

3. The specific heat of lead is $0.129 \text{ J/g}^\circ\text{C}$. Find the amount of heat released when 2.4 mol of lead are cooled from 37.3°C to 22.5°C .

$$\frac{2.4 \text{ mol Pb}}{1} \times \frac{207.2 \text{ g}}{1 \text{ mol}} = 497.28 \text{ g}$$

$$q = (497.28)(0.129)(22.5-37.3)$$

$$\boxed{q = -942 \text{ J}}$$

4. How many J of energy are needed to raise the temperature of 185 mol of water from 10.55°C to 47.32°C ?

$$\frac{185 \text{ mol H}_2\text{O}}{1} \times \frac{18.0 \text{ g}}{1 \text{ mol}} = 3330 \text{ g}$$

$$q = (3330)(4.18)(47.32-10.55)$$

$$\boxed{q = 456,485 \text{ J}}$$

5. A 150 g sample of water (initially at 45.0°C) is mixed with a 200 g sample of water (initially at 84.0°C). Find the final temperature of the system.

$$(150)(4.18)(T_f - 45.0) = -(200)(4.18)(T_f - 84.0)$$

$$627T_f - 28215 = -836T_f + 70224$$

$$\boxed{T_f = 67.3^\circ\text{C}}$$

6. A 25 g sample of iron (initially at 90.0°C) is dropped into 200 g of water (initially at 20.0°C). The final temperature of the system is 40.22°C . Find the specific heat of iron.

$$-(25)(C)(40.22-90.0) = (200)(4.18)(40.22-20.0)$$

$$1899.5C = 8543.92$$

$$\boxed{C = 0.450 \text{ J/g}^\circ\text{C}}$$

7. A 440 g sample of mercury (specific heat = $0.14 \text{ J/g}^\circ\text{C}$, initial temperature of 22.00°C) is placed into 134 g of water (initial temperature of 35.00°C). Find the final temperature of the system.

$$(440)(0.14)(T_f - 22.0) = -(134)(4.18)(T_f - 35.0)$$

$$61.6T_f - 1355.2 = -560.12T_f + 19604.2$$

$$\boxed{T_f = 33.7^\circ\text{C}}$$

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WJ Hussar



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