

Heat and Calorimetry worksheet

Specific Heat Capacity, $c_{\text{H}_2\text{O}}$ is $4.18 \text{ J/g } ^\circ\text{C}$

1. How many joules are needed to warm 25.5 grams of water from 14°C to 22.5°C ? (ans. $9.1 \times 10^2 \text{ J}$)

$$\begin{aligned} q &= m_{\text{H}_2\text{O}} c_{\text{H}_2\text{O}} \Delta T = 25.5 \text{ g} \times 4.18 \text{ J/g} \cdot ^\circ\text{C} \times (22.5 - 14)^\circ\text{C} \\ &= 906 \text{ J} = 910 \text{ J} = 9.1 \times 10^2 \text{ J} \end{aligned}$$

2. Calculate the number of joules released when 75.0 grams of water are cooled from 100.0°C to 27.5°C . (ans. $2.27 \times 10^4 \text{ J}$)

$$\begin{aligned} q &= m_{\text{H}_2\text{O}} c_{\text{H}_2\text{O}} \Delta T = 75.0 \text{ g} \times 4.18 \text{ J/g} \cdot ^\circ\text{C} \times (100.0 - 27.5)^\circ\text{C} \\ &= 22\,700 \text{ J} = 2.27 \times 10^4 \text{ J} \end{aligned}$$

3. If $1.13 \times 10^4 \text{ J}$ of heat is added to a water sample and the temperature rises from 88.0°C to its boiling point, what mass of water is in the sample? (ans. 225 g)

$$\begin{aligned} q &= m_{\text{H}_2\text{O}} c_{\text{H}_2\text{O}} \Delta T \\ 1.13 \times 10^4 \text{ J} &= m_{\text{H}_2\text{O}} \times 4.18 \text{ J/g} \cdot ^\circ\text{C} \times (100.0 - 88.0)^\circ\text{C} \\ m_{\text{H}_2\text{O}} &= \frac{1.13 \times 10^4 \text{ J}}{50.16 \text{ J/g}} = 225 \text{ g} \end{aligned}$$

4. The specific heat capacity of gold is $0.128 \text{ J/g } ^\circ\text{C}$. How much heat would be needed to warm 250.0 grams of gold from 25.0°C to 100.0°C ? (ans. $2.40 \times 10^3 \text{ J}$)

$$\begin{aligned} q &= m_{\text{Au}} c_{\text{Au}} \Delta T = 250.0 \text{ g} \times 0.128 \text{ J/g} \cdot ^\circ\text{C} \times (100.0 - 25.0)^\circ\text{C} \\ &= 2\,400 \text{ J} = 2.40 \times 10^3 \text{ J} \end{aligned}$$

5. The specific heat capacity of zinc is $0.386 \text{ J/g } ^\circ\text{C}$. How many joules would be released when 454 grams of zinc at 96.0°C were cooled to 28.0°C ? (ans. $1.19 \times 10^4 \text{ J}$)

$$\begin{aligned} q &= m_{\text{Zn}} c_{\text{Zn}} \Delta T = 454 \text{ g} \times 0.386 \text{ J/g} \cdot ^\circ\text{C} \times (96.0 - 28.0)^\circ\text{C} \\ &= 11\,900 \text{ J} = 1.19 \times 10^4 \text{ J} \end{aligned}$$

6. How much heat is absorbed by $2.50 \times 10^2 \text{ g}$ of water and the 450. g Al can (c_{Al} is $0.900 \text{ J/g } ^\circ\text{C}$) when they are heated from 10.0°C to 85.0°C ? (ans. $1.09 \times 10^5 \text{ J}$)

$$\begin{aligned} q &= m_{\text{H}_2\text{O}} c_{\text{H}_2\text{O}} \Delta T + m_{\text{Al}} c_{\text{Al}} \Delta T \\ &= (m_{\text{H}_2\text{O}} c_{\text{H}_2\text{O}} + m_{\text{Al}} c_{\text{Al}}) \Delta T \\ &= (250 \text{ g} \times 4.18 \text{ J/g} \cdot ^\circ\text{C} + 450 \text{ g} \times 0.900 \text{ J/g} \cdot ^\circ\text{C}) \times (85.0 - 10.0)^\circ\text{C} \\ &= 109\,000 \text{ J} = 1.09 \times 10^5 \text{ J} \end{aligned}$$

7. What is the temperature change if $1.386 \times 10^3 \text{ J}$ is absorbed by 60.0 g of copper which has a specific heat capacity of $0.385 \text{ J/g } ^\circ\text{C}$? (ans. 60.0°C)

$$\begin{aligned} q &= m_{\text{Cu}} c_{\text{Cu}} \Delta T \\ 1.386 \times 10^3 \text{ J} &= 60.0 \text{ g} \times 0.385 \text{ J/g} \cdot ^\circ\text{C} \times \Delta T \\ \Delta T &= \frac{1.386 \times 10^3 \text{ J}}{23.1 \text{ J/}^\circ\text{C}} = 60.0^\circ\text{C} \end{aligned}$$

Chemistry Calorimetry Problems With Answers

William Alexander Ross



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