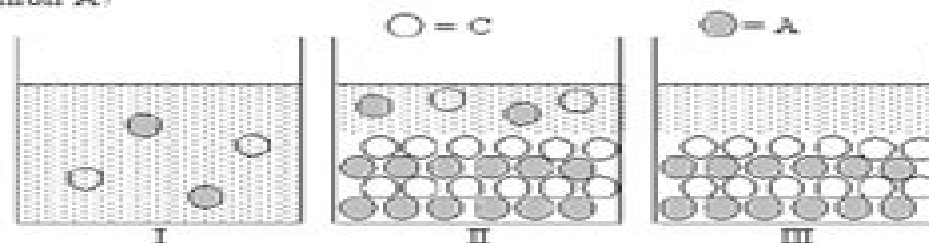


## Chapter 17 - Solubility and Complex-Ion Equilibria

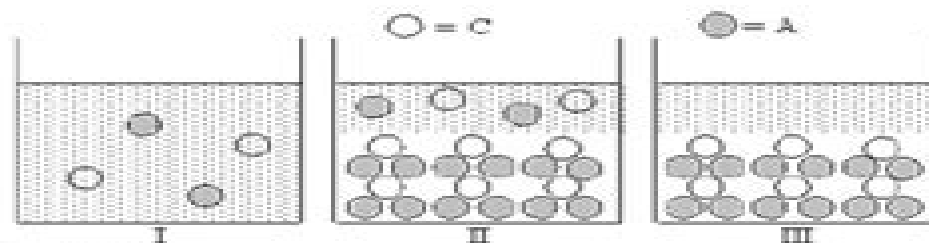
1. Cation C and anion A form an ionic compound for which  $K_{sp} = s^2$ , where  $s$  is the molar solubility of the ionic compound. Which of Figures I–III represent(s) possible results of the mixing of an aqueous solution containing cation C with an aqueous solution containing anion A?



- A) only I  
B) only III  
C) both I and III  
D) both I and II  
E) only II

ANS: D                      PTS: 1                      DIF: moderate                      REF: 17.1  
OBJ: Write solubility product expressions. (Example 17.1)  
TOP: solubility | solubility equilibria                      KEY: solubility product constant  
MSC: general chemistry

2. Cation C and anion A form an ionic compound for which  $K_{sp} = 4s^3$ , where  $s$  is the molar solubility of the ionic compound. Which of Figures I–III represent(s) possible results of the mixing of an aqueous solution containing cation C with an aqueous solution containing anion A?



- A) only III  
B) only II  
C) both I and II  
D) only I  
E) both I and III

ANS: C                      PTS: 1                      DIF: moderate                      REF: 17.1  
OBJ: Write solubility product expressions. (Example 17.1)

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**Wolfgang Guggemos**



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