

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY****The Structure of the Atom****Section 4.1 Early Theories of Matter***In your textbook, read about the philosophers, John Dalton, and defining the atom.**For each statement below, write true or false.*

- false** \_\_\_\_\_ 1. Ancient philosophers regularly performed controlled experiments.
- true** \_\_\_\_\_ 2. Philosophers formulated explanations about the nature of matter based on their own experiences.
- true** \_\_\_\_\_ 3. Both Democritus and Dalton suggested that matter is made up of atoms.
- true** \_\_\_\_\_ 4. Dalton's atomic theory stated that atoms separate, combine, or rearrange in chemical reactions.
- false** \_\_\_\_\_ 5. Dalton's atomic theory stated that matter is mostly empty space.
- false** \_\_\_\_\_ 6. Dalton was correct in thinking that atoms could not be divided into smaller particles.
- true** \_\_\_\_\_ 7. Dalton's atomic theory stated that atoms of different elements combine in simple whole-number ratios to form compounds.
- true** \_\_\_\_\_ 8. Dalton thought that all atoms of a specific element have the same mass.
- false** \_\_\_\_\_ 9. Democritus proposed that atoms are held together by chemical bonds, but no one believed him.
- true** \_\_\_\_\_ 10. Dalton's atomic theory was based on careful measurements and extensive research.
- false** \_\_\_\_\_ 11. There are no instruments powerful enough to measure atoms so that they can be seen.
- true** \_\_\_\_\_ 12. The smallest particle of an element that retains the properties of that element is called an atom.

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**CHAPTER 4 STUDY GUIDE FOR CONTENT MASTERY****Section 4.2 Subatomic Particles and the Nuclear Atom***In your textbook, read about discovering the electron and the nuclear atom.**For each item in Column A, write the letter of the matching item in Column B.*

Column A	Column B
5. _____ 1. Proposed the nuclear atomic model	a. Thomson
6. _____ 2. Determined the mass-to-charge ratio of an electron	b. Millikan
7. _____ 3. Calculated the mass of an electron	c. Rutherford

*Draw and label a diagram of each atomic model.*

4. **plum pudding model**

**Drawings should look like a ball of chocolate chip cookie dough. The chocolate chips should be labeled with negative charge or as electrons. The dough should be labeled as evenly distributed positive charges.**

5. **no clear atomic model**

**Drawings should look like a peach with a pit. The pit should be labeled nucleus and should include labeled protons and neutrons. The outer circle of the peach should be labeled electrons.**

*In your textbook, read about the discovery of protons and neutrons.**Complete the following table of protons, electrons, and neutrons characteristics.*

Particle	Symbol	Location	Relative Charge	Relative Mass
1. Proton	$p^+$	In the nucleus	$1+$	1
2. Neutron	$n^0$	In the nucleus	0	1
3. Electron	$e^-$	In the space surrounding the nucleus	$1-$	$1/1840$

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# Chemistry Change And Matter Study Guide Answers

**McGraw-Hill Staff**



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**Student Study Guide and Solutions Manual to Accompany General, Organic, and Biochemistry** Katherine J. Denniston, Katherine J. Denniston, 2006 *Chemical Principles Student's Study Guide & Solutions Manual* John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter *Basic Concepts of Chemistry, Study Guide* Leo J. Malone, 2000 The text's three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application Study Guide to accompany Basic Concepts of Chemistry, 7th Edition Leo J. Malone, 2003-02-20 Work more effectively and gauge your progress along the way This Study Guide that is designed to accompany Malone's Chemistry 7th Edition includes chapter summaries new terms self tests answers to self tests and solutions to selected problems This easy to read introduction presents chemistry as a living relevant science Chemistry 7th Edition encourages critical thinking and helps readers overcome the math difficulties that often prevent them from developing a full understanding of the subject This new seventh edition builds on its core strengths of pedagogy driving the connections between ideas mathematics in context not just an appendix and an extensive problem solving emphasis with an updated design and more molecular art In addition the seventh edition expands its applications and online options One of the briefest books in the market it still provides sufficient depth for the basic concepts of chemistry **Resources in Education**, 1984 Science Curriculum Topic Study Page Keeley, 2005-02-23 This indispensable staff development resource provides a systematic professional development strategy linking science standards and research to curriculum instruction and assessment **O Level Chemistry MCQ (Multiple Choice Questions)** Arshad Iqbal, 2019-06-27 The O

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**Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook** McGraw Hill,2001-08 Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

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